

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- | | |
|---|---|
| _____ mass | _____ percent composition. |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
| _____ 10 ²³ | |
| _____ is a description of the relative amounts of each element in a compound _____ percent composition | |
| _____ the lowest whole number ratio of the atoms of the elements in a compound _____ empirical formula | |
| _____ the volume (in milliliters) one mole of any gas at STP _____ 22.4 L | |
| _____ the smallest particle of a compound is a compound _____ molar mass | |
| _____ the number representing 6.02 x 10 ²³ representative particles of a substance _____ Avogadro's number | |
| _____ 1 mol = 6.02 ²³ particles _____ empirical formula | |
| _____ the number representing the number of atoms of an element in a molecule of a compound _____ molecular formula | |

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Completion Use the completion strategy to strengthen knowledge of the terms and your understanding of the concepts involved with the terms of the definitions completed with a term (don't guess) or explain.

Measuring whole units (counting, weighing, or measuring volume for a single quantity) called the _____ method. _____ Representative representative particles (one mole of a substance is _____ Avogadro's number _____)

6.02x10²³ is the number of particles of a compound, substance, or mixture that _____ molar mass _____ of the substance (the molar mass _____ other one calculates the molar mass for atoms (the molar mass _____ molar mass _____ of the compound). These are _____ 6.02x 10²³ _____ representative particles one mole of any substance.

True or False

_____ 1 mole of any substance is 6.02 x 10²³ particles.

_____ 1 mole of a substance contains 6.02 x 10²³ atoms.

_____ 1 mole of a substance contains 6.02 x 10²³ representative particles. **True**

_____ 1 mole of H₂O is composed of one mole of hydrogen and one mole of oxygen. **True**

_____ 1 mole mass of nitrogen gas (N₂) is 28.02 g. **True**

Problems

10 Chemical Quantities Chapter Quiz Answer Key

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A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

10 Chemical Quantities Chapter Quiz Answer Key:

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plot edit the story opens up with a young boy named danny going to a science museum he sees indians bears eskimos guns and swords finally he immediately gets drawn to the dinosaur exhibit and would be delighted to find a living dinosaur the dinosaurs in the dinosaur exhibit are really models and not real

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