Stoichiometry: Mass - Mole Worksheet

Based on the following chemical equation, please answer the following questions!

Fe (SO₄), + 6 NaCl --> 2 FeCl, + 3 Na₂SO₄

Calculate the amount of moles of NaCl needed to produce 75 g of FeCl_s.

| 9 FeCl, | 1 mole FeCl, | mole NaCl |
|---------|--------------|-----------|
| | 9 FeCl, | mole FeCl |

Calculate the amount of Fe, (SO,), in grams produce 215 g Na, SO,.

| 9 Na ₃ SO ₄ | 1 mole Na SO | mole Fe ₂ (SO ₄), | 9 Fe ₂ (SO ₄), |
|-----------------------------------|-----------------------------------|--|--|
| | 9 Na ₂ SO ₄ | mole Na,SO, | 1 mole Fe ₂ (SO ₄), |

Calculate the amount in grams of FeCl, produced with 95g of Fe (SO₄),

| 9 Fe ₂ (SO ₄), | mole Fe2(SO4), | mole FeCl, | 9 FeCl, |
|---------------------------------------|----------------|---------------|--------------|
| | 9 Fe.(SO,), | mole Fe(SO,), | 1 mole FeCl, |

Calculate the amount of moles of Fe₂(SO₄) needed to react with 123.5 g of NaCl

| 9 NaCl | 1 mole NaCl | mole Fe(SO ₄), | |
|--------|-------------|----------------------------|--|
| | 9 NaCl | mole NaCl | |

Based on the following chemical equation, please answer the following questions!

| 9 Ca(HCO ₃) ₂ | 1 mole Ca(HCO ₂), | mole CaCO, |
|--------------------------------------|-------------------------------|---------------|
| | 9 Ca(HCO ₁), | mole Ca(HCO,) |

Answers For Study Guide Mass And Moles

Douglas Brandt

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