

অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



পটীক্ষার কপূর, এখানে সেরে পড়িষ্ঠা, আরও কলেজের শিক্ষার্থীর কলেজসমূহের পটীক্ষা এবং মতামত প্রদানের ক্ষমতা
পূর্ণাঙ্গ উত্তর অধ্যয়নিক সেবা প্রদান। এগুলোর অনুশীলন করলে তুমি এ অধ্যয়ন থেকে সেরে পড়িষ্ঠা প্রদান করবে।
এদের উত্তর লিখতে পারবে সহজেই।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১০/১০

- মোলের প্রকণ কী?
- "কেলস পানিদ্রুত লবণের প্রাকৃত্য, তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।
- A-পাত্রে বিদ্যমান এসিডের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা প্রক্রিয়াকৃত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL দ্রবণে কোন দ্রবের 1 mol দ্রবীভূত থাকলে উক্ত দ্রবকে মোলার দ্রবণ বলে।

কেলস পানিদ্রুত লবণের প্রাকৃত্য তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে প্রাকৃত্য বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কেলস পানি দ্রুত লবণ নিষ্কৃতি হয়। তখন প্রাকৃত্য হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পানিতে দ্রবীভূত করার সময় বিয়োজন প্রক্রিয়াটি তাপশক্তি বিক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে প্রাকৃত্য হ্রাস ঘটে।

১

আমরা জানি,
 $w = MV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$

অতএব, দ্রবীভূত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদীপকের A-পাত্রে অবস্থিত,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$

দ্রবণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
দ্রবণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

দ্রবীভূত H_2SO_4 এর ভর, $w = ?$

পাত্র-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

পাত্র-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্বন্ধিত বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে

0.077 mol H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে $0.048 \text{ mol H}_2\text{SO}_4$ অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লীয়।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর ত্রুটি সহ পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{w}{v}$ HCl এর নমুনা তৈরি করে।

সি. কে. ১০/১০/১০

- নোড কাকে বলে?
- Cl_2 অপোলার— ব্যাখ্যা করো।
- উদীপকের এসিডটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সমান হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন গঠিত হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও H এর তড়িৎ ঋণাত্মকতা 2.1 তাই Cl_2 বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌলিক অণুর ক্ষেত্রে এহুণ ঘটে থাকে।

উদীপকে বর্ণিত এসিডটি হলো 2% $\left(\frac{w}{v}\right) \text{ HCl}$

2% $\left(\frac{w}{v}\right) \text{ HCl}$ এর অর্থ হলো—

100 mL দ্রবণে HCl এর ভর 2g

$\therefore 1000 \text{ mL}$ " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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2018 Edit, Mark Dixon



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