

100. Carbon tetrachloride consists of one carbon and four chlorine atoms. How many atoms of carbon are there in 1.00 mole of CCl_4 molecules?

$$\text{one mole of } \text{CCl}_4 \text{ contains } 1 \text{ mole of } \text{C} \text{ atoms} \quad 1 \text{ mole of } \text{C} \text{ atoms} = 6.02 \times 10^{23} \text{ atoms}$$

$$1 \text{ mole of } \text{CCl}_4 \text{ contains } 6.02 \times 10^{23} \text{ atoms of C}$$

$$\text{answer} = 6.02 \times 10^{23} \text{ atoms} \quad 100$$

101. There are 3 monatomic gases in equal molar amounts. What are they?

101. Helium, neon, and argon.

Helium, neon, and argon are the only monatomic gases that are gases at room temperature.

Helium, neon, and argon are the only monatomic gases that are gases at room temperature.

102. Suggest other three gas law relationships between pressure and volume.

Boyle's law, Charles' law, and Gay-Lussac's law.

103. Sodium chloride is a compound of sodium and chlorine. How many moles of sodium are there in 1.00 mole of NaCl ?

1.00 mole of NaCl contains 1.00 mole of Na.

What is the molar mass of sodium? 22.99 g/mol

$$\text{Na} = 22.99 \text{ g/mol}$$

$$= 22.99 \text{ g/mol} \times 1 \text{ mole} = 22.99 \text{ g}$$



Answer To F321 Chemistry May Series

Mike Smith



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