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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	kJ/mol	kJ/mol
N-H	408	391
O-H	463	463

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an **N-O bond** in kJ/mol .

Use the following steps.

- Calculate the **energy needed to break bonds in the reactants**.

$$(408 \times 12) + 463 \times 4$$

$$408 \times 12 = 4896$$

$$463 \times 4 = 1852$$

$$4896 + 1852 = 6748 \text{ kJ/mol}$$
- Calculate the **energy released when bonds form in the products**.

$$2 \times (391 \times 4 + 463 \times 2) = 2 \times (1564 + 926) = 2 \times 2490 = 4980 \text{ kJ/mol}$$
- Calculate the energy of **an N-O bond**.

$$6748 - 4980 = 1768 \text{ kJ/mol}$$

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Christian G. Meyer



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