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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-H
Bond energy in kJ/mol	436	435

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an N-N bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.
 $(4H_2) + 1(N_2)$
 $4(436) + 1(945)$

3350 kJ
- Calculate the energy released when bonds form in the products.
 $2(2NH_3) = 2(3 \times N-H) = -90$
 $3350 - 90 = -90$

3260 kJ
- Calculate the energy of an N-N bond.
 $6 \times (N-H) = 2676$
 $N-N = \frac{3260 - 2676}{1}$

584 kJ/mol

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Zhe Xu



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