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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	O-H
Bond energy in kJ/mol	436	463

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an **O-H bond** in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(4H_2) + 3(O_2)$$

$$4(436) + 3(498)$$

$$3150 \text{ kJ}$$
- Calculate the energy released when bonds form in the products.

$$\text{Bond broken} - \text{Bond formed} = -90$$

$$3150 - x = -90$$

$$x = 3240$$

$$3240 \text{ kJ}$$
- Calculate the energy of **an O-H bond**.

$$6 \times \text{O-H} = 3240$$

$$\text{O-H} = \frac{3240}{6}$$

$$540 \text{ kJ/mol}$$

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Christian G. Meyer



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