Name: Key

For each reaction:

- \* Assign Oxidation Numbers
- \* Write an Oxidation and Reduction Half Reaction
- \*Circle the Oxidizing Agent and Underline the Reducing Agent.

1) Fe<sup>-3</sup> + 5n → Fe<sup>-2</sup> + 5n<sup>-2</sup>

Oxidation: Snº -> Snº + 2e

Reduction: (Fe') . Is' -> Fe'

2) AlF, \* Bo -> Al \* BoF.

Oxidation: Bot + Bort + 2e-

Reduction: (Al-3) + 3e - Al-

3) H<sub>2</sub>O<sub>2</sub> + H<sub>2</sub>SO<sub>2</sub> → H<sub>2</sub>O + H<sub>2</sub>SO<sub>2</sub>

Oxidation: 5-1 + 5-1 + 2e

Reduction: (O1) + Ie: + O-

4)  $H_1SO_1 + KMnO_4 \rightarrow MnSO_4 + K_1SO_4 + H_2O_5$ 

Oxidation: 5" + 5" + 2e

Reduction: (Mn-) + 5e - Mn-2

5) K2Cr2O7 + H2C2O4 + H2SO4 + K2SO4 + Cr2(SO4)3 + CO2 + H2O

Oxidation: C1 -> C-1 + 1e

Reduction: (Cris) . 3c -> Cris

# **Analyzing Oxidation Reduction Reactions Answers Sheet**

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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures NBS Special Publication ,1968

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