

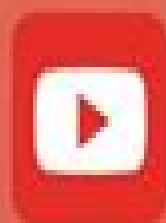
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0620 CHEMISTRY

MAY/JUNE - 2024

PAPER - 61

0620_S24_QP_61 (Q3-4)



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Table 1 shows the tests and the student's observations for solid P.

Tests	Observations
Test 1 Heat about half of solid P in boiling tube with test tube in fumes of HCl .	<u>disappears + effs</u> the white solid forms a colourless fume, steam effervesces out from the boiling tube and condensation is seen at the top of the boiling tube. after a while the colourless fume becomes a white solid.
Test 2 The remaining solid P is dissolved in water to form solution P. Solution P is divided into three portions. To the first portion of solution P, a few drops of NaOH are added, and the mixture is stirred thoroughly and then is heated. Repeat the procedure and then using red litmus paper at the mouth of the boiling tube.	<u>no effs</u> <u>no H_2S gas</u> <u>no H_2SO_4</u> <u>no change</u> The fumes red litmus paper remains red.
Test 3 To the second portion of solution P, a few drops of dilute H_2SO_4 are added, and a few drops of BaCl_2 are added.	<u>white precipitate</u> <u>SO_4^{2-}</u>
Test 4 To the third portion of solution P, with a few drops of acidified aqueous potassium manganate(VII).	the solution becomes light purple <u>SO_3^{2-}</u> <u>purple $\rightarrow$ fading</u>

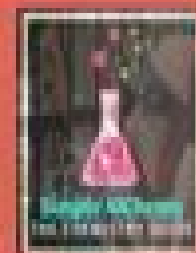
30. The observations in test 1 show that solid P is hydrated.

31. Suggest a chemical test to show that the condensation at the top of the boiling tube contains water.

— anhydrous blue cobalt chloride paper

— pink $\rightarrow$ eff

— anhydrous white crumbly solid ✓



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R Barnett

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