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0620 CHEMISTRY

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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-H
Bond energy in kJ/mol	436	391

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of an O-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.
 $(40H) + 1(N_2)$
 $4(436) + 1(945)$

3350 kJ
- Calculate the energy released when bonds form in the products.
 $6(N-H) = 6(391) = -2346$
 $3350 - 2346 = -996$

-996 kJ
- Calculate the energy of an O-H bond.
 $6 \times (N-H) = 2346$
 $N-H = \frac{2346}{6}$

391 kJ/mol

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A Gutmann



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