

Worksheet 1.6

Supplemental Instruction
Iowa State University

Leader: Lea

Course: Chem 177

Instructor: Dr. Schewe-Miller

Date: 2/5/2013

- 1.) A piece of aluminum foil 1.00 cm square and 0.550 mm thick is allowed to react with bromine to form aluminum bromide.

a.) How many moles of aluminum were used? (density Al = 2.699 g/cm³)

$$.550\text{cm} \times \frac{1\text{cm}}{10\text{mm}} = .055\text{cm}$$

$$1\text{cm} \times .055\text{cm} = .055\text{cm}^2 \times \frac{2.699\text{g}}{\text{cm}^3} = .148445\text{g Al} \times \frac{1\text{mol Al}}{26.98\text{g Al}} = \boxed{5.50 \times 10^{-3}\text{mol Al}}$$

b.) How many grams of aluminum bromide form, assuming the aluminum reacts completely? What type of reaction is this?



$$5.50 \times 10^{-3}\text{mol Al} \times \frac{2\text{mol AlBr}_3}{2\text{mol Al}} \times \frac{266.69\text{g AlBr}_3}{1\text{mol AlBr}_3} = \boxed{1.47\text{g}}$$

composition reaction

2.) The chemical formula for benzene is C₆H₆.

a. Determine the percent composition of H and C for benzene.

$$6(12.01) + 6(1.01) = 78.12$$

$$\text{C: } \frac{72.06}{78.12} \times 100 = 92\% \quad \text{H: } 100 - 92\% = 8\%$$

$$\boxed{\begin{array}{l} \text{C: } 92\% \\ \text{H: } 8\% \end{array}}$$

b. Write a balanced combustion reaction for benzene.



c. Assuming you have 364g of benzene, how many moles of O₂ will be required to completely combust this amount of benzene?

$$364\text{g} \times \frac{1\text{mol C}_6\text{H}_6}{78.12\text{g C}_6\text{H}_6} \times \frac{15\text{mol O}_2}{2\text{mol C}_6\text{H}_6} = \boxed{35.0\text{mol}}$$

d. How many grams of CO₂ will this produce? How many grams of H₂O?

$$\text{CO}_2: 364\text{g C}_6\text{H}_6 \times \frac{1\text{mol C}_6\text{H}_6}{78.12\text{g C}_6\text{H}_6} \times \frac{12\text{mol CO}_2}{2\text{mol C}_6\text{H}_6} \times \frac{44.01\text{g CO}_2}{1\text{mol CO}_2} = \boxed{1.23 \times 10^3\text{g CO}_2}$$

$$\text{H}_2\text{O}: 364\text{g C}_6\text{H}_6 \times \frac{1\text{mol C}_6\text{H}_6}{78.12\text{g C}_6\text{H}_6} \times \frac{6\text{mol H}_2\text{O}}{2\text{mol C}_6\text{H}_6} \times \frac{18\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = \boxed{2.52 \times 10^3\text{g H}_2\text{O}}$$

e. You perform the combustion, and 956g of CO₂ are produced. What is the percent yield?

$$\frac{\text{Actual}}{\text{Theoretical}} \times 100 = \frac{956\text{g CO}_2}{1230\text{g CO}_2} \times 100 = \boxed{77.7\%}$$

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Laurie Sorge



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Comprehensive Chemistry XI Dr. B. Kapila,S. K. Khanna,2010-11 Comprehensive chemistry according to the new syllabus prescribed by Central Board of Secondary Education CBSE Essential Laboratory Mathematics Catherine W. Johnson,Daniel L. Timmons,Pamela E. Hall,2009-12-03 This hands on manual with pedagogical features that draw the learner into the content offers clear and complete coverage of the mathematical topics most often used in today s clinical and medical laboratories Furthermore it provides a solid foundation for subsequent courses in the laboratory sciences The first two chapters present a review of basic mathematical concepts The remainder of the book provides students with a realistic means to build on previously learned concepts both mathematical and scientific to refine their mathematical skills and to gauge their mastery of those skills Outstanding features Each chapter opens with an outline objectives and key terms Key terms highlighted within the text are listed and defined in the glossary Margin problems and practice problem sets provide the chance to gain immediate proficiency Laboratory exercises and review problems allow students to apply what they ve learned and assess their understanding and progress A special calculator icon signals explanations of calculator use for a

particular mathematical function Study hints Keys to Success offer practical suggestions and guidance for maximizing achievement The workbook design enables users to solve problems and take notes directly on the pages Prentice Hall Chemistry ,2000

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