

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- | | |
|-------------------------|--|
| _____ mass | _____ percent composition. |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ standard temperature and pressure (STP, 1 atm) |
| _____ STP | |

1. _____ is a description of the relative amounts of each element in a compound.
2. _____ composition
3. The lowest whole number ratio of the atoms of the elements in a compound is the _____ formula.
4. The lowest structural formula and mass of any gas at STP is _____.
5. The lowest structural formula of a compound is _____.
6. The lowest structural formula of a compound is _____.
7. The lowest structural formula of a compound is _____.
8. The lowest structural formula of a compound is _____.
9. The lowest structural formula of a compound is _____.
10. The lowest structural formula of a compound is _____.

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Complete the following statement by choosing the correct word or phrase. Consider the following concepts: mass, volume, density, and molar mass.

- Choosing whole units of counting, or mass, and volume for a single quantity called the _____ mass. _____ represents the number of particles in a sample of a substance is _____ Avogadro's number.
- The number of particles in a compound, which is used together with _____ percent composition, of the substance gives the molar mass. _____ represents the number of particles in a sample of a substance.

Practice Problems

1. The mass of a sample of a compound is 10.0 g.

2. The mass of a sample of a compound is 10.0 g.

3. The mass of a sample of a compound is 10.0 g.

4. The mass of a sample of a compound is 10.0 g.

5. The mass of a sample of a compound is 10.0 g.

Problems

10 Chemical Quantities Chapter Quiz Answer Key

Arshad Iqbal



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Chemical Surety United States. Department of the Army,1995

Chemistry: The Central Science Theodore L. Brown,H. Eugene LeMay Jr.,Bruce E. Bursten,Catherine Murphy,Patrick Woodward,Steven Langford,Dalius Sagatys,Adrian George,2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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