

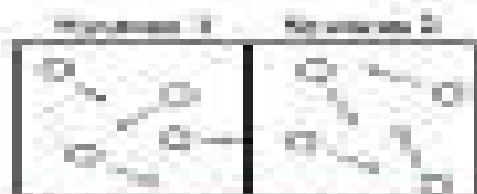
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J . The first mixing of particles and molecules results in change, instantly increasing the 100 J temperature. However, this thermally distribution curve shows that the average velocity is still for the distribution's maximum value to change. The average temperature is constant.

3. Sketch a curve used to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more spread out than in System 2, indicating a lower average kinetic energy. Therefore, System 1 is at a lower temperature than System 2.
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.
 Thermal energy will be transferred from System 1 to System 2. System 1 is at a lower temperature than System 2, meaning the particles in System 1 have a lower average kinetic energy. When the two systems are in contact, the particles in System 1 will collide with the particles in System 2, transferring energy to them. This process will continue until the two systems reach thermal equilibrium.

Answer Key Unit 6 Thermochemistry

Steven S. Zumdahl



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