

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing without performing stoichiometric calculations.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of a solid reagent that would be produced.
5. To compare the theoretical amount to the actual amount of a solid reagent and calculate the percent yield.

Reagents and Equipment:

Reagent 1: Solid Chemical Reagent 1: 0.10 mol

Reagent 2: Solid Chemical Reagent 2: 0.10 mol

Reagent 3: Solid Reagent 3: 0.10 g

Reagent 4: Solid Reagent 4: 0.10 g

Reagent 5: Solid Reagent 5: 0.10 g

Reagent 6: Solid Reagent 6: 0.10 g

Reagent 7: Solid Reagent 7: 0.10 g

Procedure:

Reagents and Calculations:

1. Calculations:

$$\text{Moles of reagent 1} = \frac{\text{Mass of reagent 1}}{\text{Molar mass of reagent 1}} = \frac{0.10 \text{ g}}{100.0 \text{ g/mol}} = 0.0010 \text{ mol}$$

$$\text{Moles of reagent 2} = \frac{\text{Mass of reagent 2}}{\text{Molar mass of reagent 2}} = \frac{0.10 \text{ g}}{100.0 \text{ g/mol}} = 0.0010 \text{ mol}$$

$$\text{Moles of reagent 3} = \frac{\text{Mass of reagent 3}}{\text{Molar mass of reagent 3}} = \frac{0.10 \text{ g}}{100.0 \text{ g/mol}} = 0.0010 \text{ mol}$$

$$\text{Moles of reagent 4} = \frac{\text{Mass of reagent 4}}{\text{Molar mass of reagent 4}} = \frac{0.10 \text{ g}}{100.0 \text{ g/mol}} = 0.0010 \text{ mol}$$

2. Chemical Equations:

3. Reagents:

$$\text{Reagent 1} + \text{Reagent 2} \rightarrow \text{Reagent 3} + \text{Reagent 4} + \text{Reagent 5} + \text{Reagent 6} + \text{Reagent 7}$$

$$\text{Reagent 1} + \text{Reagent 2} \rightarrow \text{Reagent 3} + \text{Reagent 4} + \text{Reagent 5} + \text{Reagent 6} + \text{Reagent 7}$$

$$\text{Reagent 1} + \text{Reagent 2} \rightarrow \text{Reagent 3} + \text{Reagent 4} + \text{Reagent 5} + \text{Reagent 6} + \text{Reagent 7}$$

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