Ionic bond is more solid where covalent bond has a lower melting point, and ionic bonds result from transfer of electrons, whereas covalent bonds are formed by sharing

the bonding atoms obtain stable electron configurations; since the stable configuration is usually eight electrons in the outermost shell

15. A shared pair of electrons is called a **bonding pair**, while a pair that is associated with only one atom—and therefore not involved in bonding—is called a **lone pair**.

24. Atoms and molecules . CO2

29.

a. molecular

b. ionia

c. ionic

d. molecular

33.

a. 3 Mg. 2 P. 8 O

b. 1 Ba , 2 Cl

c. 1 Fe, 2 N, 4 O

d. 1 Ca. 2 O. 2 H

39.



43.

a. SrSe

b. BaCl2

c. Na2S

d. Al2O3

45. The product of the charges of the two ions is one factor of lattice energy. CsF has a +1 ion charge product, whereas BaO has a +4 ion charge product. The lattice energy is four times bigger because the product is four times greater.

57.

a. NaHSO3

b. LiMnO4

c. AgNO3

d. K2SO4

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Danielle Baeyens-Volant, Pascal Laurent, Nathalie Warzée

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