

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive the empirical formula of a pure substance with percent composition (molecular formula) information.
2. To calculate the number of moles of a solid from weighing information on the substance.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction Information:

Reaction of Sodium Chloride Solution: 0.10 mol

Silver Nitrate Solution: 0.10 mol

AgNO₃ 10% Silver Paper: 0.10 g

AgNO₃ 10% Silver Paper + Dry Solid: 0.10 g

Reaction of Sodium Chloride Solution: 0.10 mol

Reaction of Silver Nitrate Solution: 0.10 mol

Reaction of Silver Nitrate Solution: 0.10 mol

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{NaCl (aq)} + \text{AgNO}_3 \text{ (aq)} \rightarrow \text{AgCl (s)} + \text{NaNO}_3 \text{ (aq)}$$

$$\text{NaCl (aq)} + \text{AgNO}_3 \text{ (aq)} \rightarrow \text{AgCl (s)} + \text{NaNO}_3 \text{ (aq)}$$

$$\text{NaCl (aq)} + \text{AgNO}_3 \text{ (aq)} \rightarrow \text{AgCl (s)} + \text{NaNO}_3 \text{ (aq)}$$

$$\text{NaCl (aq)} + \text{AgNO}_3 \text{ (aq)} \rightarrow \text{AgCl (s)} + \text{NaNO}_3 \text{ (aq)}$$

2. Reaction:

Reaction:

$$\text{NaCl (aq)} + \text{AgNO}_3 \text{ (aq)} \rightarrow \text{AgCl (s)} + \text{NaNO}_3 \text{ (aq)}$$

$$\text{NaCl (aq)} + \text{AgNO}_3 \text{ (aq)} \rightarrow \text{AgCl (s)} + \text{NaNO}_3 \text{ (aq)}$$

$$\text{NaCl (aq)} + \text{AgNO}_3 \text{ (aq)} \rightarrow \text{AgCl (s)} + \text{NaNO}_3 \text{ (aq)}$$

Answers For Lab 6b Heath Chemistry

A Loxley



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