

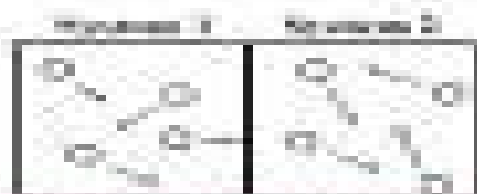
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J . The first mixing of particles and molecules results in change, instantly increasing the 100 J to 150 J (temperature increased). The Boltzmann distribution curve shows that the average velocity is still for the distribution's maximum (area under the curve) is 150 J (temperature increased).

3. Sketch a curve used to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more clustered together, indicating a lower average KE. In System 2, the particles are more spread out, indicating a higher average KE. The particles in System 1 are more clustered together, indicating a lower average KE. The particles in System 2 are more spread out, indicating a higher average KE.
- b. An initial amount of kinetic energy for System 1. Justify your answer in terms of molecular collisions.
 All particles in System 1 are at a lower temperature than System 2. The particles in System 1 are more clustered together, indicating a lower average KE. The particles in System 2 are more spread out, indicating a higher average KE. The particles in System 1 are more clustered together, indicating a lower average KE. The particles in System 2 are more spread out, indicating a higher average KE.

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