

## Chemistry Unit 8 Worksheet 4

### Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$  (s), that must decompose to produce potassium chloride,  $\text{KCl}$  (s), and 1.8 moles of oxygen gas.

|           |                                                           |      |      |                                     |
|-----------|-----------------------------------------------------------|------|------|-------------------------------------|
| Equation: | $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ |      |      | 1.2 moles<br>$\text{KClO}_3$ needed |
| Before:   | 1.2                                                       | 0    | 0    |                                     |
| Change:   | -1.2                                                      | +1.2 | +1.8 |                                     |
| After:    | 0                                                         | 1.2  | 1.8  |                                     |

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

|           |                                                                   |          |         |       |                                    |
|-----------|-------------------------------------------------------------------|----------|---------|-------|------------------------------------|
| Equation: | $\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ |          |         |       | 0.200 moles<br>$\text{HCl}$ needed |
| Before:   | 0.100                                                             | ?        | 0       | 0     |                                    |
| Change:   | -0.100                                                            | -0.200 + | 0.100 + | 0.100 |                                    |
| After:    | 0                                                                 | ?        | 0.100   | 0.100 |                                    |

3. Ethane,  $\text{C}_2\text{H}_6$  reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

|           |                                                                                          |        |        |        |
|-----------|------------------------------------------------------------------------------------------|--------|--------|--------|
| Equation: | $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$ |        |        |        |
| Before:   | 2.20                                                                                     | xs     | 0      | 0      |
| Change:   | -2.20                                                                                    | - 7.70 | + 4.40 | + 6.60 |
| After:    | 0                                                                                        | xs     | 4.40   | 6.60   |

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

|           |                                                                                                 |          |         |         |
|-----------|-------------------------------------------------------------------------------------------------|----------|---------|---------|
| Equation: | $\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$ |          |         |         |
| Before:   | 0.0040                                                                                          | XS       | 0       | 0       |
| Change:   | - 0.0040                                                                                        | - 0.0080 | +0.0040 | +0.0080 |
| After:    | 0                                                                                               | XS       | 0.0040  | 0.0080  |

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

# Answer Sheet Chemistry Unit 8 Worksheet 4

**Kenneth T. Henson**



## **Answer Sheet Chemistry Unit 8 Worksheet 4:**

## Unveiling the Magic of Words: A Review of "**Answer Sheet Chemistry Unit 8 Worksheet 4**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Answer Sheet Chemistry Unit 8 Worksheet 4**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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