

Question Number	Acceptable Answers	Reject	Mark																									
23(a)	<table><tr><th>Element</th><th>%</th><th></th><th>mol</th><th>Ratio</th></tr><tr><td>C</td><td>40.44</td><td>÷ 12</td><td>= 3.37</td><td>2.999</td></tr><tr><td>H</td><td>7.87</td><td>÷ 1</td><td>= 7.87</td><td>7.004</td></tr><tr><td>O</td><td>35.96</td><td>÷ 16</td><td>= 2.2475</td><td>2.000</td></tr><tr><td>N</td><td>15.73</td><td>÷ 14</td><td>= 1.12357</td><td>1.000</td></tr></table> (1) (1) Empirical formula = C₃H₇O₂N (1) Symbols in any order Stand alone mark No TE on incorrect ratio I IGNORE significant figure and rounding errors except 1sf in mole calculation	Element	%		mol	Ratio	C	40.44	÷ 12	= 3.37	2.999	H	7.87	÷ 1	= 7.87	7.004	O	35.96	÷ 16	= 2.2475	2.000	N	15.73	÷ 14	= 1.12357	1.000		3
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N	15.73	÷ 14	= 1.12357	1.000																								

Question Number	Acceptable Answers	Reject	Mark
23(b)(i)	Peak at $m/e = 89$ labelled M^+ ALLOW Any clear label e.g. $\text{C}_3\text{H}_7\text{O}_2\text{N}^+$		1

Question Number	Acceptable Answers	Reject	Mark
23(b)(ii)	$M_r = 89 = M_r(\text{C}_3\text{H}_7\text{O}_2\text{N})$ so molecular formula is $\text{C}_3\text{H}_7\text{O}_2\text{N}$ Symbols in any order IGNORE structural and displayed formulae	Answer with no explanation $M_r = 90$ $\text{C}_3\text{H}_7\text{O}_2\text{N}^+$	1

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L Towne



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