# 11-2 Practice Problems

 Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:

 $Li_3N + 3H_2O \rightarrow NH_3 + 3LiOH$   $| Mail LiN \times \frac{3Mail LiOH}{1001 LiOH} \times \frac{34a LiOH}{1001 LiOH} = 0.78a$ 

 What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?

029 NoI + Cl2 - 2NoCI + Iz 029 NoI + Cl2 - 2NoCI + Iz 029 NoI x 1mol x 2mol NoI (mol NoC) NoCI

> Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:

0-85gCyH10+1302-8002+10H20 1 × 1mol × 8mol (0) × 144g = 2.58g 2 mol CyH10 1 mol CO z

 Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:

0.4695bD= x [mol x 2mol 5b, 0, x 121.769 fa.389 1 291.529 [mol 5b, 0, x 120.769 5b]

What mass of hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>) must decompose to produce 0.77 g of water?
2H<sub>2</sub>O<sub>2</sub> → 2H<sub>2</sub>O + O<sub>2</sub>

0.779H20 x 1mol x 2H20 x 349 = 1.4594202

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?

0.69gCoz x 1mol x 2mol Co x 289 = 0.449 1 2mol Co x 1mol = 0.449  Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:

 $\frac{N8(NO_3)_2 + 2N8OH \rightarrow N8(OH)_2 + 2N8NO_3}{O.73_3 V(NO_3)_2} \times \frac{1 mot}{185_3} \times \frac{2 mat}{1 mot} \frac{N8(NO_3)_2}{1 mot} \times \frac{25_3}{1 mot} \frac{(0.68)_3}{N8NO_3}$ 

 Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:

 $0.64_0H_2O \times \frac{CaC_2 + 2H_2O - Ca(OH)_2 + C_2H_2}{189} \times \frac{Imol(a(0H)_2 + C_2H_2)}{Imol} \times \frac{749}{Imol} = 1.32_9$   $Ca(OH)_2$ 

 How many grams of ozone (O<sub>3</sub>) must decompose to produce 0.87 g of oxygen?

0.8790 x 1mg/ x 2mg/ x 48903 = 0.87903

- Find the mass of sugar (C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation: C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> → 2C<sub>2</sub>H<sub>6</sub>O + 2CO<sub>2</sub>
- How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP? The balanced equation is S + O<sub>2</sub> → SO<sub>2</sub>.
- Find the mass of benzene (C<sub>6</sub>H<sub>6</sub>) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:

2C<sub>6</sub>H<sub>6</sub> + 15O<sub>2</sub> → 6H<sub>2</sub>O + 12CO<sub>2</sub>

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Ying Liu

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