

1 From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014
(Question 1).

a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	1.6725×10^{-24}	1.6725×10^{-24}	0.0009×10^{-24}

- i An atom of hydrogen can be represented as ${}^1\text{H}$. Use data from Table 1.1 to calculate the mass of this hydrogen atom.

$$1.6734 \times 10^{-27} \text{ kg}$$

(1 mark)

- ii Which one of the following is a fundamental particle that would not be deflected by an electric field?

- A electron
B neutron
C proton

Write the correct letter, A, B, or C, on the line below.

B

(1 mark)

- b A naturally occurring sample of the element boron has a relative atomic mass of 10.8. In this sample, boron exists as two isotopes, ${}^{10}\text{B}$ and ${}^{11}\text{B}$.

- i Calculate the percentage abundance of ${}^{10}\text{B}$ in this naturally occurring sample of boron.

$$10.8 = (10)(x) + (11)(1-x)$$

$$10.8 = 10x + 11 - 11x$$

$$10.8 - 11 = 10x - 11x$$

$$-0.2 = -x$$

$$x = 0.2 \quad 0.2 \times 100 = 20\%$$

(2 marks)

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Jim Breithaupt, Catherine Jones



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web to fill out the kone fault code list follow these steps 01 begin by gathering all relevant fault codes from the kone system 02 ensure that each fault code is properly labeled