- From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014 (Question 1).
- a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	1.6725 × 10⁻³⁴	1.6725 × 10 <sup>-24</sup>	0.0009 × 10 <sup>-24</sup>

An atom of hydrogen can be represented as 'H. Use data from Table 1.1 to calculate the mass of this hydrogen atom.

1.6734x10<sup>-27</sup> kg

(1 mark)

Which one of the following is a fundamental particle that would not be deflected by an electric field?

A electron

B neutron

C proton

Write the correct letter, A, B, or C, on the line below.

8

- (fimark)
- b A naturally occurring sample of the element boron has a relaive atomic mass of 10.8. In this sample, boron exists as two isotopes, ™B and ™B.
  - Calculate the percentage abundance of <sup>10</sup>B in this naturally occurring sample of boron.

$$10.8 = (10)(x) + (11)(1-x)$$

$$10.8 = 10x + 11 - 11x$$

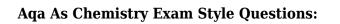
$$10.8 - 11 = 10x - 11x$$

$$-0.2 = -\infty$$

(2 marks)

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L Darling-Hammond



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