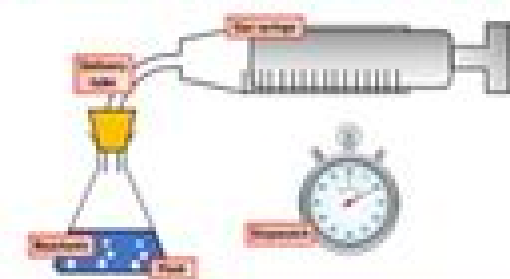


Calculating Rate

$$\text{Rate} = \frac{\text{amount of product made}}{\text{Time taken}}$$

Investigating Rate of Reaction

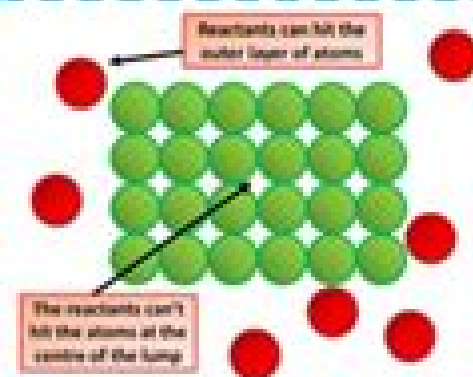
1. Support a gas syringe with a stand, boss and clamp.
2. Using a measuring cylinder, add 10 cm³ of dilute hydrochloric acid to a conical flask.
3. Add 0.4 g of calcium carbonate to the flask, immediately connect the gas syringe and start a stop clock.
4. Every 10 seconds, record the volume of gas produced.
5. When the reaction is complete, clean the apparatus as directed by a teacher.
6. Repeat steps 1 to 5 with different concentrations of hydrochloric acid.



The Effect of Surface Area

For a given mass of a solid, large lumps have smaller surface area to volume ratios than smaller lumps or powders. If a large lump is divided or ground into a powder:

- its total volume stays the same
- the area of exposed surface increases
- the surface area to volume ratio increases



The Effect of Temperature

The greater the frequency of successful collisions, the greater the rate of reaction, if the temperature of the reaction mixture is increased:

- reactant particles move more quickly
- the energy of the particles increases
- the frequency of successful collisions between reactant particles increases
- the proportion of collisions which are successful increases
- the rate of reaction increases

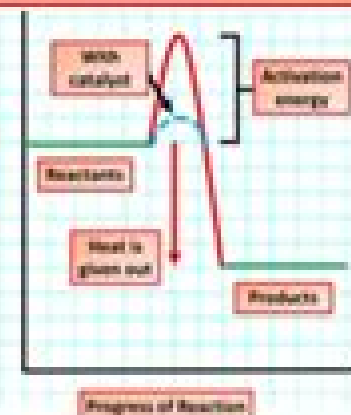
The Effect of a Catalyst

A catalyst is a substance that:

- speeds up the rate of a reaction
- does not alter the products of the reaction
- is not chemically changed or used up at the end of the reaction

Only a very small mass of catalyst is needed to increase the rate of a reaction. However, not all reactions are the same. Different catalysts are used for different reactions.

Energy

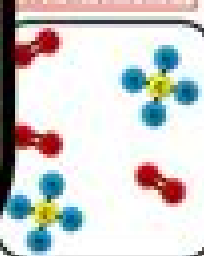


AQA GCSE Chemistry Rate of Reaction Knowledge Organiser

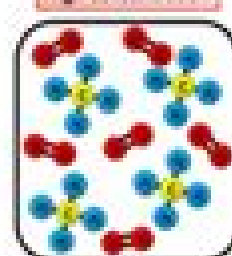
Reversible reaction where the reactants form products that, in turn, form the reactants back. Reversible reactions will reach an equilibrium where the concentrations of the reactants and products will no longer change.

Concentration

Low concentration

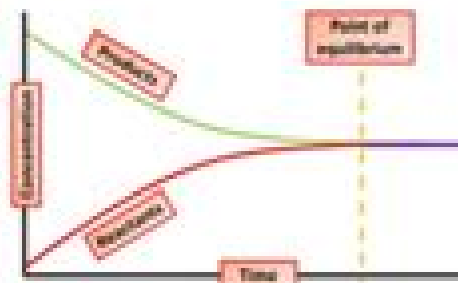


High concentration



Equilibrium

Chemical equilibrium is the point at which the concentrations of reactants and products do not change with time. It appears as though the reaction has stopped but in fact the rates of the forward and reverse reactions are equal so reactants and products are being created at the same rate.



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Table of Contents Aqa Gcse Chemistry

1. Understanding the eBook Aqa Gcse Chemistry
 - The Rise of Digital Reading Aqa Gcse Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Aqa Gcse Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aqa Gcse Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aqa Gcse Chemistry
 - Personalized Recommendations
 - Aqa Gcse Chemistry User Reviews and Ratings
 - Aqa Gcse Chemistry and Bestseller Lists
5. Accessing Aqa Gcse Chemistry Free and Paid eBooks

- Aqa Gcse Chemistry Public Domain eBooks
- Aqa Gcse Chemistry eBook Subscription Services
- Aqa Gcse Chemistry Budget-Friendly Options
- 6. Navigating Aqa Gcse Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Aqa Gcse Chemistry Compatibility with Devices
 - Aqa Gcse Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aqa Gcse Chemistry
 - Highlighting and Note-Taking Aqa Gcse Chemistry
 - Interactive Elements Aqa Gcse Chemistry
- 8. Staying Engaged with Aqa Gcse Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aqa Gcse Chemistry
- 9. Balancing eBooks and Physical Books Aqa Gcse Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aqa Gcse Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Aqa Gcse Chemistry
 - Setting Reading Goals Aqa Gcse Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Aqa Gcse Chemistry
 - Fact-Checking eBook Content of Aqa Gcse Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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