0 1 . 3

A second series of experiments was carried out to investigate how the rate of the reaction varies with temperature.

The results were used to obtain a value for the activation energy of the reaction, E.

Identical amounts of reagents were mixed at different temperatures.

The time taken, t, for a fixed amount of bromine to be formed was measured at different temperatures.

The results are shown in Table 3.

Table 3

Temperature, T / K	1/K-1	Time, t	1/5-1	In 1/
286	3.50 × 10 <sup>-3</sup>	54	1.85 × 10 <sup>-2</sup>	-3.99
295	3.39 × 10 <sup>-3</sup>	27	3.70 × 10 <sup>-2</sup>	
302		15	6.67 × 10 <sup>-2</sup>	-2.71
312	3.21 × 10 <sup>-3</sup>	8	1.25 × 10 <sup>-1</sup>	-2.08

Complete Table 3.

[2 marks]

0 1 . 4

The Arrhenius equation can be written as

$$\ln k = -\frac{E_a}{R} \left( \frac{1}{T} \right) + C_1$$

In this experiment, the rate constant, k, is directly proportional to  $\frac{1}{t}$ 

Therefore

$$\ln \frac{1}{t} = -\frac{E_a}{R} \left( \frac{1}{T} \right) + C_2$$

where C<sub>1</sub> and C<sub>2</sub> are constants.

Use values from **Table 3** to plot a graph of  $\ln \frac{1}{t}$  (y axis) against  $\frac{1}{\tau}$  on the grid.

Use your graph to calculate a value for the activation energy, in kJ mol<sup>-1</sup>, for this reaction.

The value of the gas constant, R = 8.31 J K-1 mol-1

[6 marks]



# **Aqa P2 2013 May 23**

**Robert Besancon** 

#### Aga P2 2013 May 23:

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