

1 From AQA Chemistry Unit 1 Foundation Chemistry CHEM1 May 2014
(Question 1).

a Table 1.1 shows some data about fundamental particles in an atom.

Table 1.1 Mass of fundamental particles

Particle	proton	neutron	electron
Mass / g	1.6725×10^{-24}	1.6725×10^{-24}	0.0009×10^{-24}

- i** An atom of hydrogen can be represented as ${}^1\text{H}$. Use data from Table 1.1 to calculate the mass of this hydrogen atom.

$$1.6734 \times 10^{-27} \text{ kg}$$

(1 mark)

- ii** Which one of the following is a fundamental particle that would not be deflected by an electric field?

- A** electron
B neutron
C proton

Write the correct letter, **A**, **B**, or **C**, on the line below.

B

(1 mark)

- b** A naturally occurring sample of the element boron has a relative atomic mass of 10.8. In this sample, boron exists as two isotopes, ${}^{10}\text{B}$ and ${}^{11}\text{B}$.

- i** Calculate the percentage abundance of ${}^{10}\text{B}$ in this naturally occurring sample of boron.

$$10.8 = (10)(x) + (11)(1-x)$$

$$10.8 = 10x + 11 - 11x$$

$$10.8 - 11 = 10x - 11x$$

$$-0.2 = -x$$

$$x = 0.2 \quad 0.2 \times 100 = 20\%$$

(2 marks)

Aqa Chemistry Exam Style Questions

Nora Henry, Alyn G. McFarland



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Table of Contents Aqa Chemistry Exam Style Questions

1. Understanding the eBook Aqa Chemistry Exam Style Questions
 - The Rise of Digital Reading Aqa Chemistry Exam Style Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Aqa Chemistry Exam Style Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aqa Chemistry Exam Style Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aqa Chemistry Exam Style Questions

- Personalized Recommendations
- Aqa Chemistry Exam Style Questions User Reviews and Ratings
- Aqa Chemistry Exam Style Questions and Bestseller Lists
- 5. Accessing Aqa Chemistry Exam Style Questions Free and Paid eBooks
 - Aqa Chemistry Exam Style Questions Public Domain eBooks
 - Aqa Chemistry Exam Style Questions eBook Subscription Services
 - Aqa Chemistry Exam Style Questions Budget-Friendly Options
- 6. Navigating Aqa Chemistry Exam Style Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Aqa Chemistry Exam Style Questions Compatibility with Devices
 - Aqa Chemistry Exam Style Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aqa Chemistry Exam Style Questions
 - Highlighting and Note-Taking Aqa Chemistry Exam Style Questions
 - Interactive Elements Aqa Chemistry Exam Style Questions
- 8. Staying Engaged with Aqa Chemistry Exam Style Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aqa Chemistry Exam Style Questions
- 9. Balancing eBooks and Physical Books Aqa Chemistry Exam Style Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aqa Chemistry Exam Style Questions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Aqa Chemistry Exam Style Questions
 - Setting Reading Goals Aqa Chemistry Exam Style Questions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Aqa Chemistry Exam Style Questions

- Fact-Checking eBook Content of Aqa Chemistry Exam Style Questions
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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