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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = \frac{\ln \left(\frac{N_0}{N} \right)}{\ln 2}$ $n = \frac{\ln \left(\frac{125}{5} \right)}{\ln 2} = 4.32$ 125 g 5 what is the half life of an isotope that is 80 remained after 16 days solution remaining 80 100 therefore $n = \frac{\ln \left(\frac{100}{80} \right)}{\ln 2} = 1.68$ now using the half life equation in example 1 we have

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