

CHEMICAL BONDING

↳ Force of attraction present between any chemical species like, atoms, ions & molecules.

Bond formation → Exothermic

Bond dissociation → Endothermic

Chemical bonds on the basis of Bond Energy

Strong bond

- ① Ionic bond
- ② Covalent bond
- ③ co-ordinate bond
- ④ Metallic bond

Weak bond

- ① Hydrogen bond
- ② Vander waal's bond

Cause of chemical combination

Modern concept :- Every atom has a tendency to form Max^m possible bonds in order to minimize its potential energy & gain stability.

Classical concept (octet rule) :-

↳ Every atom has a tendency to achieve stable outer octet ($8e^-$ in valence shell).

Different theories to explain bonding

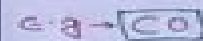
- ① Octet rule
- ② VSEPR (valence shell electron pair repulsion theory)
- ③ VBT (valence bond theory)
- ④ MOT (Molecular orbital theory)

Kossel-Lewis Approach (octet rule):

- Based on inertness of noble gas
- Atoms can be combine either by transfer or by sharing of valence electrons in order to have an octet ($8e^-$ in their valence shells).
- In case of H & He duplet ($2e^-$ in valence shell) will be completed.

How to write Lewis dot structure & bond formation of molecules.

$$\text{No. of Bonds} = \frac{(\text{No. of atoms} \times \text{Val. e}^-) - \text{Available val. e}^-}{2}$$



$$\text{No. of bonds} = \frac{(2 \times 2) - (4 + 6)}{2} = 3$$

btw C & O



$$\text{No. of bonds} = \frac{(2 \times 3) - (6 \times 3)}{2}$$



$$= 3$$



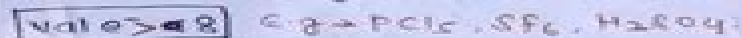
$$\text{No. of bonds} = \frac{2 \times 2 - (5 \times 3) + 1}{2}$$



$$= 4.$$

Limitations of octet rule

(i) Existence of Hypervalent compound.



(ii) Existence of Hypovalent compound.



(iii) Odd electron species do not follow octet rule. e.g. → NO, NO_2

(iv) Existence of compounds of inert gases with oxygen & fluorine like $\text{XeF}_2, \text{KrF}_2, \text{XeOF}_2$ etc.

(v) does not explain the relative stability of the molecules being totally silent about the energy of a molecule.

(vi) This theory does not account for the shape of molecules.

⇒ formal charge :-

F.C = Itne hone chahiye - Itne hain on atom (val. e⁻ in free state) (val. e⁻ in given structure)

$$\text{F.C} = \text{Total val. e}^- - [2 \times \text{l.p} + \text{Bond pair}]$$

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Herbert Beall, John Trimbur



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