

PHYSICS

QUESTIONS

1. According to the Bohr theory, which of the following transitions will give rise to the most energetic photon?

- (1) $n=3 \rightarrow n=2$ (2) $n=4 \rightarrow n=3$
(3) $n=4 \rightarrow n=2$ (4) $n=3 \rightarrow n=1$

2. A car has a mass of 75 kg and speed of 20 m/s. If the speed has to be measured with the accuracy of 1%, then the uncertainty in its position is

- (1) 1.0×10^{-14} m (2) 1.0×10^{-16} m
(3) 1.0×10^{-18} m (4) 1.0×10^{-20} m

3. The ground state of hydrogen atom is at -13.6 eV. The value of the first excited state of hydrogen atom is

- (1) -3.4 eV (2) -13.6 eV
(3) -10.2 eV (4) -1.5 eV

4. The sample of a pure, first number of excited lines obtained during transition of electron from 4th excited state to ground state is

- (1) 6 (2) 4
(3) 10 (4) 8

5. The mass of proton having wavelength 1.5×10^{-10} m is

- (1) 1.6×10^{-27} kg (2) 1.6×10^{-29} kg
(3) 1.6×10^{-31} kg (4) 1.6×10^{-33} kg

6. The value of $\left(\frac{h^2}{8ma^2}\right)$ is maximum for

- (1) a constant (2) a variable
(3) a small (4) a large

7. Which of the following pair are correct?

- (1) a and b

8. Maximum kinetic energy of photoelectrons emitted when a light of frequency 1.5×10^{15} Hz is incident on a metal surface whose threshold frequency is equal to 1.0×10^{15} Hz, is 5.0×10^{-19} J

- (1) 5.0×10^{-19} J (2) 1.0×10^{-19} J
(3) 5.1×10^{-19} J (4) 6.0×10^{-19} J

9. The number of photons of light having wavelength 660 nm which can provide 100 J of energy is equal to 1.6×10^{21}

- (1) 1.2×10^{21} (2) 2.2×10^{21}
(3) 3.2×10^{21} (4) 4.2×10^{21}

10. Which of the following arrangement shows the correct order of maximum multiplicity?

- (1) $\uparrow \uparrow \uparrow \uparrow \uparrow$ (2) $\uparrow \uparrow \uparrow \uparrow \uparrow$
(3) $\uparrow \uparrow \uparrow \uparrow \uparrow$ (4) $\uparrow \uparrow \uparrow \uparrow \uparrow$

11. Which of the following is correct?

- (1) $1s$ orbital is lower in energy than $2s$ orbital
(2) $2s$ orbital is lower in energy than $2p$ orbital
(3) $3s$ orbital is lower in energy than $3p$ orbital
(4) $4s$ orbital is lower in energy than $4p$ orbital

12. Which of the following is correct?

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(4) $4s$ orbital is lower in energy than $4p$ orbital

- (1) $1s$ and $2s$ (2) $2s$ and $3s$
(3) $3s$ and $4s$ (4) $4s$ and $5s$

Space for Rough Work

PHYSICS

QUESTIONS

1. Which of the following pair of orbitals will have maximum overlap along the x-axis?

- (1) $1s$ and $2s$
(2) $2p_x$ and $3p_x$
(3) $2s$ and $3s$
(4) $2p_x$ and $3p_y$

2. Maximum wavelength of spectral series of Balmer series for hydrogen atom is 656.3 nm. The wavelength of the spectral series is

- (1) 656.3 nm (2) 656.3 nm
(3) 656.3 nm (4) 656.3 nm

3. Which of the following is correct?

- (1) $1s$ and $2s$ orbitals have same energy
(2) $2s$ and $3s$ orbitals have same energy
(3) The energy of $2s$ orbital is less than that of $2p$ orbital
(4) Angular momentum of an electron in $2s$ and $2p$ orbitals is same
(5) The ionization enthalpy of $2s$ and $2p$ orbitals is same
(6) The ionization enthalpy of $2s$ and $2p$ orbitals is same

- (1) $1s$ and $2s$
(2) $2s$ and $3s$
(3) $2s$ and $2p$
(4) $2p$ and $3p$

[BIOLOGY]

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(3) $2s$ and $2p$
(4) $2p$ and $3p$

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Lei Huang



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