

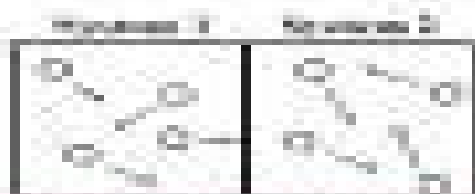
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles that move fast enough, the more kinetic energy we will see. Temperature is a scalar. The Boltzmann probability curves show that the average velocity we will see for the distribution is increasing (less to large) as the temperature increases.

3. Sketch what you expect to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the gas particles are more spread out, so all the individual particles have the same mass, depending on their velocities KE is $\frac{1}{2}mv^2$, so all the particles have the same mass, velocity is the more important factor affects KE. In the diagram in System 1 it is wider, which particles are moving at more velocities and it represents a lower average KE.
- b. Are there different total kinetic energy for System 2? Justify your answer in terms of molecular collisions.
KE can be transferred from System 2 to System 1 if the particles are going faster, depending on the system of (same) temperature. Temperature represents the gas particles are not necessarily that much as the spread of molecular velocity, but velocity is calculated by mass of particles KE. The molecular velocity will predict what is going to happen KE and KE.

Answer Key Unit 6 Thermochemistry

Arshad Iqbal



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