

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing without performing stoichiometric calculations.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of a solid reagent that would be produced.
5. To compare the theoretical amount to the actual amount of a solid reagent and calculate the percent yield.

Reagents and Equipment:

Reagent 1: Sodium Chloride (NaCl) 25.0 g

Reagent 2: Calcium Chloride (CaCl₂) 25.0 g

Reagent 3: Silver Nitrate (AgNO₃) 5.00 g

Reagent 4: Silver Nitrate (AgNO₃) 5.00 g

Reagent 5: Silver Nitrate (AgNO₃) 5.00 g

Reagent 6: Silver Nitrate (AgNO₃) 5.00 g

Reagent 7: Silver Nitrate (AgNO₃) 5.00 g

Procedure:

Reagents and Calculations:

1. Calculations:

$$\text{moles of sodium chloride} = \frac{\text{mass of NaCl}}{\text{molar mass of NaCl}} = \frac{25.0 \text{ g}}{58.44 \text{ g/mol}} = 0.428 \text{ mol}$$

$$\text{moles of calcium chloride} = \frac{\text{mass of CaCl}_2}{\text{molar mass of CaCl}_2} = \frac{25.0 \text{ g}}{110.98 \text{ g/mol}} = 0.225 \text{ mol}$$

$$\text{moles of silver nitrate produced} = 0.428 \text{ mol} + 0.225 \text{ mol} = 0.653 \text{ mol}$$

$$\text{mass of silver nitrate produced} = 0.653 \text{ mol} \times \frac{169.87 \text{ g/mol}}{1} = 110.7 \text{ g}$$

2. Chemical Equations: $\text{NaCl} + \text{CaCl}_2 \rightarrow \text{CaCl}_2 + \text{NaCl}$

3. Silver Nitrate:

$$\text{Silver: } 1 \text{ mol AgNO}_3 \rightarrow 1 \text{ mol Ag} \quad \text{Silver: } 1 \text{ mol AgNO}_3 \rightarrow 1 \text{ mol Ag} \quad \text{Silver: } 1 \text{ mol AgNO}_3 \rightarrow 1 \text{ mol Ag}$$

$$\text{Silver: } 1 \text{ mol AgNO}_3 \rightarrow 1 \text{ mol Ag} \quad \text{Silver: } 1 \text{ mol AgNO}_3 \rightarrow 1 \text{ mol Ag}$$

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Siddappa N.Byrareddy



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