

16. Increasing the temperature generally.....
- decreases the solubility of a solid in a liquid but increases the solubility of a gas in a liquid.
 - increases the solubility of a solid in a liquid but decreases the solubility of a gas in a liquid.
 - increases the solubility of both a solid and a gas in a liquid.
 - decreases the solubility of both a solid and a gas in a liquid.
17. A white precipitate was formed when $\text{BaCl}_{2(aq)}$ was added to an aqueous solution of a salt X. The precipitate dissolved in dilute HCl with rapid effervescence. Salt X is likely to contain
- SO_4^{2-} ions.
 - SO_3^{2-} ions.
 - NO_3^- ions.
 - CO_3^{2-} ions.
18. Before a reaction could take place, there should be
- ionization of reactant particles.
 - breakage of bonds of products.
 - breakage of bonds of reactants.
 - ionization of product particles.
19. Consider the following reaction equation:
- $$\text{CaO} + \text{SiO}_2 \longrightarrow \text{CaSiO}_3$$
- Silicon(IV) oxide is acting as
- a basic oxide.
 - a reducing agent.
 - an acidic oxide.
 - an oxidizing agent.
20. Which of the following acids would form normal salt only?
- Tetraoxosulphate(VI) acid
 - Trioxonitrate(V) acid
 - Tetraoxophosphate(V) acid
 - Trioxosulphate(IV) acid
21. What is the partial pressure of oxygen at s.t.p. in a gaseous mixture containing 100 cm^3 of oxygen and 900 cm^3 of nitrogen gas?
- 1.0 atm
 - 0.1 atm
 - 9.0 atm
 - 0.9 atm
22. Graphite is used as a dry lubricant due to the presence of
- mobile electrons.
 - free electrons.
 - octahedral structures.
 - layered structures.
23. Which of the following gases is alkaline?
- NO_2
 - CO_2
 - NH_3
 - N_2O
24. Which of the following statements about an equilibrium system is correct?
- Forward and backward reactions occur at the same rate.
 - The concentrations of reactants must equal that of the products.
 - The concentrations of reactants and products can be changed by adding a catalyst.
 - The concentrations of reactants and products are not affected by a change in temperature.

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