

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Sub-levels have different amounts of energy. For example, orbitals closer to the nucleus have more energy than orbitals further from the nucleus. The following orbitals have different levels of energy:



For each pair, write sublevel(s) for which orbitals are included. Then draw orbital(s) representing an orbital.



Now draw 4 orbitals representing the set of 4p orbitals. Write down the names of each orbital. Remember, the 4d orbitals have higher energy than 4s orbitals and the 4f orbitals have a higher energy than 4d orbitals.

When drawing energy levels, try to represent the amount of energy possible. (This is called the **Aufbau Principle**). An electron will never be added before the sublevel is filling and no electron will be added until after the sublevel is filled. Additionally, no electron will be added until (This is called the **Mutual Exclusion Principle**). When two electrons occupy the same orbital they must spin in opposite directions – one clockwise and the other anticlockwise.

Critical Thinking Questions

- a) How many electrons would it take to fill a sublevel before beginning to be added? **Sublevels become filled, all four orbitals are filled before electrons will be added to the 3p.**
b) What is the first element before enough electrons are there to fill 3 orbitals? **After you provide this:**
The 1st 3d element will be Scandium. The element is Scandium 21.
- a) How many electrons would it take to fill a sublevel before beginning to be added? **Sublevels become full, all four orbitals are filled before the 3p, 4p, 5p, 6p.**
b) What is the first element before enough electrons are being placed electrons into 4 orbitals? **The 21st element will be Scandium. The element is Scandium 21.**

Information: Hund's Rule

Electrons prefer to be "paired" or "coupled". Sublevels closer to the nucleus have higher energy. Paired electrons are the lowest energy. For example, four electrons are paired in one orbital. Hund's orbital diagram is shown below.



If a sublevel is represented as four orbitals, electrons will go into each orbital diagram. The first two electrons will be paired in one orbital. The other two electrons will be paired in one orbital – orbitals in the 2p sublevel will go.



Hund's rule will be followed if the electrons are paired. Hund's rule states that electrons have a choice of occupying several orbitals before the orbitals are filled. A sublevel is filled when all orbitals are filled. Hund's rule states that electrons will occupy orbitals. Therefore, Choice A is correct. Hund's rule states that electrons will occupy orbitals before they are paired in one orbital.

The following are the electron orbital diagrams for the elements, nitrogen and oxygen. Write the element's 3p sublevel and explain to show Hund's rule. The 3p sublevel is filled before the 4p sublevel is filled.



Critical Thinking Questions

- How many "paired" electrons are in a sublevel? **1**
- Why does carbon's orbitals diagram look like this? Why not 3p orbitals? **Why not 3p orbitals? The 3p orbitals are already full. Hund's rule states that electrons will occupy orbitals before they are paired in one orbital.**
- Why do elements like nitrogen and oxygen have 3p orbitals? **1s, 2s, 2p, 3s, 3p, 4s, 4p, 5s, 5p, 6s, 6p, 7s, 7p, 8s, 8p, 9s, 9p, 10s, 10p, 11s, 11p, 12s, 12p, 13s, 13p, 14s, 14p, 15s, 15p, 16s, 16p, 17s, 17p, 18s, 18p, 19s, 19p, 20s, 20p, 21s, 21p, 22s, 22p, 23s, 23p, 24s, 24p, 25s, 25p, 26s, 26p, 27s, 27p, 28s, 28p, 29s, 29p, 30s, 30p, 31s, 31p, 32s, 32p, 33s, 33p, 34s, 34p, 35s, 35p, 36s, 36p, 37s, 37p, 38s, 38p, 39s, 39p, 40s, 40p, 41s, 41p, 42s, 42p, 43s, 43p, 44s, 44p, 45s, 45p, 46s, 46p, 47s, 47p, 48s, 48p, 49s, 49p, 50s, 50p, 51s, 51p, 52s, 52p, 53s, 53p, 54s, 54p, 55s, 55p, 56s, 56p, 57s, 57p, 58s, 58p, 59s, 59p, 60s, 60p, 61s, 61p, 62s, 62p, 63s, 63p, 64s, 64p, 65s, 65p, 66s, 66p, 67s, 67p, 68s, 68p, 69s, 69p, 70s, 70p, 71s, 71p, 72s, 72p, 73s, 73p, 74s, 74p, 75s, 75p, 76s, 76p, 77s, 77p, 78s, 78p, 79s, 79p, 80s, 80p, 81s, 81p, 82s, 82p, 83s, 83p, 84s, 84p, 85s, 85p, 86s, 86p, 87s, 87p, 88s, 88p, 89s, 89p, 90s, 90p, 91s, 91p, 92s, 92p, 93s, 93p, 94s, 94p, 95s, 95p, 96s, 96p, 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