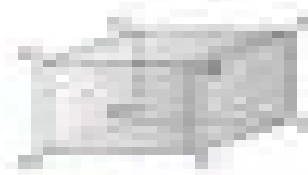


Answer following



- 1. Substrate and temperature: 20°C 30°C 40°C
- 2. A graph of initial velocity: $\text{mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $\text{mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $\text{mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$
- 3. A graph of initial velocity: $\text{mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $\text{mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ $\text{mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$

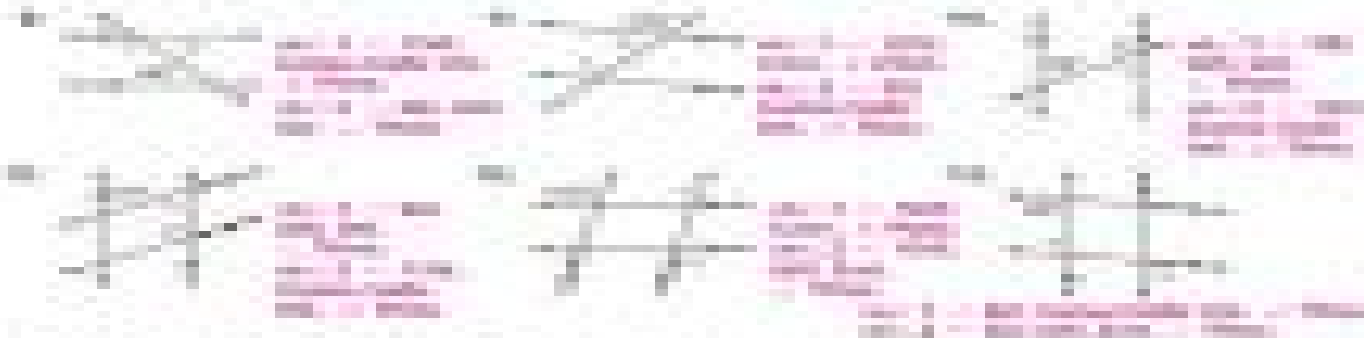
Calculate all values of K_m and V_{max} for the following reactions:
 (Note: K_m and V_{max} are given in $\text{mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ and $\text{mg} \cdot \text{min}^{-1} \cdot \text{mg}^{-1}$ respectively)

- 1. $K_m = 10$ $V_{max} = 10$
- 2. $K_m = 10$ $V_{max} = 10$
- 3. $K_m = 10$ $V_{max} = 10$
- 4. $K_m = 10$ $V_{max} = 10$



Answer 10/10

Graphs of initial velocity vs. substrate concentration for the following reactions:



Graphs of initial velocity vs. substrate concentration for the following reactions:



Graphs of initial velocity vs. substrate concentration for the following reactions:

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M Mosston



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