

Chapter 20 Worksheets: Redox ANSWERS

1. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

- $2\text{Sr} + \text{O}_2 \longrightarrow 2\text{SrO}$ $\text{Sr}^0 \text{ to } \text{Sr}^{2+}$; oxidized/reducing agent $\text{O}^0 \text{ to } \text{O}^{2-}$; reduced/ox. ag.
- $2\text{Li} + \text{S} \longrightarrow \text{Li}_2\text{S}$ $\text{Li}^0 \text{ to } \text{Li}^{+}$; oxidized/ox. ag. $\text{S}^0 \text{ to } \text{S}^{2-}$; reduced/ox. ag.
- $2\text{Cu} + \text{Br}_2 \longrightarrow 2\text{CuBr}$ $\text{Cu}^0 \text{ to } \text{Cu}^{+}$; oxidized/ox. ag. $\text{Br}^0 \text{ to } \text{Br}^{-}$; reduced/ox. ag.
- $3\text{Mg} + \text{N}_2 \longrightarrow \text{Mg}_3\text{N}_2$ $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$; oxidized/ox. ag. $\text{N}^0 \text{ to } \text{N}^{3-}$; reduced/ox. ag.
- $4\text{Fe} + 3\text{O}_2 \longrightarrow 2\text{Fe}_2\text{O}_3$ $\text{Fe}^0 \text{ to } \text{Fe}^{3+}$; oxidized/ox. ag. $\text{O}^0 \text{ to } \text{O}^{2-}$; reduced/ox. ag.
- $\text{Cl}_2 + 2\text{NaBr} \longrightarrow 2\text{NaCl} + \text{Br}_2$ $\text{Cl}^0 \text{ to } \text{Cl}^{-}$; reduced/ox. ag. $\text{Br}^{-} \text{ to } \text{Br}^0$; oxidized/ox. ag.
- $\text{Si} + 2\text{F}_2 \longrightarrow \text{SiF}_4$ $\text{Si}^0 \text{ to } \text{Si}^{4+}$; oxidized/ox. ag. $\text{F}^0 \text{ to } \text{F}^{-}$; reduced/ox. ag.
- $\text{Mg} + 2\text{HCl} \longrightarrow \text{MgCl}_2 + \text{H}_2$ $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$; oxidized/ox. ag. $\text{H}^{+} \text{ to } \text{H}^0$; reduced/a.
- $2\text{Na} + 2\text{H}_2\text{O} \longrightarrow 2\text{NaOH} + \text{H}_2$ $\text{Na}^0 \text{ to } \text{Na}^{+}$; oxidized/a. $\text{H}^{+} \text{ to } \text{H}^0$; reduced/a.

11. Give the oxidation number of each kind of atom or ion.

- | | | | | | | |
|------------|-------|--------------------|---------------------|---------------------|------------|-------------|
| a. sulfate | b. Fe | c. S^{2-} | d. Fe^{3+} | e. Sn^{4+} | f. nitrate | g. ammonium |
| 2- | 0 | 2- | 3+ | 4+ | 1- | 1+ |

12. Calculate the oxidation number of chromium in each of the following.

- | | | | | |
|----------------------------|---------------------------------------|--------------------|-------------|---------------|
| a. Cr_2O_3 | b. $\text{Na}_2\text{Cr}_2\text{O}_7$ | c. CrSO_4 | d. chromate | e. dichromate |
| 3+ | 6+ | 2+ | 7+ | 6+ |

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)

- $\text{C} + \text{H}_2\text{SO}_4 \longrightarrow \text{CO}_2 + \text{SO}_2 + \text{H}_2\text{O}$ $\text{C}^0 \text{ to } \text{C}^{4+}$; oxidized $\text{S}^{6+} \text{ to } \text{S}^{4+}$; reduced
- $\text{HNO}_3 + \text{HI} \longrightarrow \text{NO} + \text{I}_2 + \text{H}_2\text{O}$ $\text{N}^{5+} \text{ to } \text{N}^{2+}$; reduced $\text{I}^{-} \text{ to } \text{I}^0$; oxidized
- $\text{KMnO}_4 + \text{HCl} \longrightarrow \text{MnCl}_2 + \text{Cl}_2 + \text{H}_2\text{O} + \text{KCl}$ $\text{Mn}^{7+} \text{ to } \text{Mn}^{2+}$; reduced $\text{Cl}^{-} \text{ to } \text{Cl}^0$; oxidized
- $\text{Sn} + \text{HNO}_3 \longrightarrow \text{Sn}_2\text{O}_3 + \text{NO} + \text{H}_2\text{O}$ $\text{Sn}^0 \text{ to } \text{Sn}^{2+}$; oxidized $\text{N}^{5+} \text{ to } \text{N}^{2+}$; red.

14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

- | | |
|-------------------------------|--------------------------|
| a. oxidizing agent: sulfur | reducing agent: carbon |
| b. oxidizing agent: nitrogen | reducing agent: iodine |
| c. oxidizing agent: manganese | reducing agent: chlorine |
| d. oxidizing agent: nitrogen | reducing agent: antimony |

Chapter 20 Oxidation Reduction Reactions Worksheet

Arshad Iqbal



Chapter 20 Oxidation Reduction Reactions Worksheet:

Marine Geochemistry Roy Chester, Tim D. Jickells, 2012-08-24 Marine Geochemistry offers a fully comprehensive and integrated treatment of the chemistry of the oceans their sediments and biota The first edition of the book received strong critical acclaim and was described as a standard text for years to come This third edition of Marine Geochemistry has been written at a time when the role of the oceans in the Earth System is becoming increasingly apparent Following the successful format adopted previously this new edition treats the oceans as a unified entity and addresses the question how do the oceans work as a chemical system To address this question the text has been updated to cover recent advances in our understanding of topics such as the carbon chemistry of the oceans nutrient cycling and its effect on marine chemistry the acidification of sea water and the role of the oceans in climate change In addition the importance of shelf seas in oceanic cycles has been re evaluated in the light of new research Marine Geochemistry offers both undergraduate and graduate students and research workers an integrated approach to one of the most important reservoirs in the Earth System Additional resources for this book can be found at www.wiley.com/go/chester/marinegeochemistry *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Educart One-shot Science CBSE Class 10 Question Bank 2025-26 on new Syllabus 2026 (Strictly for Boards Exam)* Educart, 2025-05-26 Book Structure Handpicked Important Ch wise Q s How Good is the Educart One shot Question Bank Covers essential topics with concise yet detailed explanations to help you grasp concepts quickly Aligned with the latest rationalised syllabus to ensure relevant and up to date content Includes a variety of High Order Thinking Questions to build problem solving skills Step by step answers to NCERT and exemplar problems for better understanding Previous Year DIKSHA Platform Questions to give you real exam exposure Smart Study Tips Tricks to strengthen your conceptual clarity and boost confidence Why choose this book Get the Educart One Shot Question Bank today and take your exam preparation to the next level **Books in Print Supplement**, 2002

MnM_POW-Science-PM-10 (Updated) Vibha Arora, Anju Sachdeva, Sushma Sardana, MnM_POW Science PM 10 Updated *Index Medicus*, 2004 Vols for 1963 include as pt 2 of the Jan issue Medical subject headings **Oxidation-reduction Reactions** Irish Science Teachers' Association, 1983 *Principles of Redox Reactions* Malika Ammam, 2018-03-05 As a teacher of physical chemistry I noticed that students even in advanced classes have difficulties in understanding the basics of redox chemistry In this Section 3 I attempted to discuss some fundamental principles related to redox processes by focusing on the species that might lose or gain electrons determination of the oxidation numbers or states of atoms in compounds and ways of balancing redox reactions To further clarify the discussed concepts numerous questions and problems with detailed answers are provided Most of these questions are formulated by students like you I believe that this Section 3 would greatly help students with levels varying from high school to advanced university classes **Oxidation-reduction Reactions** Irish Science Teachers' Association. Chemistry Syllabus Committee, 1988* **A Level Chemistry Multiple Choice Questions**

and Answers (MCQs) Arshad Iqbal, 2020-04-10 Previously published as A Level Chemistry MCQs Multiple Choice Questions and Answers Quiz Tests with Answer Keys by Arshad Iqbal A Level Chemistry Multiple Choice Questions and Answers MCQs A Level Chemistry quizzes practice tests with answer key provides mock tests for competitive exams to solve 1745 MCQs A Level Chemistry MCQs helps with theoretical conceptual and analytical study for self assessment career tests This book can help to learn and practice A Level Chemistry quizzes as a quick study guide for placement test preparation A level Chemistry Multiple Choice Questions and Answers MCQs is a revision guide with a collection of trivia quiz questions and answers on topics Alcohols and esters atomic structure and theory benzene chemical compound carbonyl compounds carboxylic acids acyl compounds chemical bonding chemistry of life electrode potential electrons in atoms enthalpy change equilibrium group IV groups II and VII halogenoalkanes hydrocarbons introduction to organic chemistry ionic equilibria lattice energy moles and equations nitrogen and sulfur organic and nitrogen compounds periodicity polymerization rates of reaction reaction kinetics redox reactions and electrolysis states of matter transition elements to enhance teaching and learning A level Chemistry Quiz Questions and Answers also covers the syllabus of many competitive papers for admission exams of different universities from chemistry textbooks on chapters Alcohols and Esters Multiple Choice Questions 27 MCQs Atomic Structure and Theory Multiple Choice Questions 37 MCQs Benzene Chemical Compound Multiple Choice Questions 41 MCQs Carbonyl Compounds Multiple Choice Questions 29 MCQs Carboxylic Acids and Acyl Compounds Multiple Choice Questions 27 MCQs Chemical Bonding Multiple Choice Questions 213 MCQs Chemistry of Life Multiple Choice Questions 29 MCQs Electrode Potential Multiple Choice Questions 62 MCQs Electrons in Atoms Multiple Choice Questions 53 MCQs Enthalpy Change Multiple Choice Questions 45 MCQs Equilibrium Multiple Choice Questions 50 MCQs Group IV Multiple Choice Questions 53 MCQs Groups II and VII Multiple Choice Questions 180 MCQs Halogenoalkanes Multiple Choice Questions 33 MCQs Hydrocarbons Multiple Choice Questions 53 MCQs Introduction to Organic Chemistry Multiple Choice Questions 52 MCQs Ionic Equilibria Multiple Choice Questions 56 MCQs Lattice Energy Multiple Choice Questions 33 MCQs Moles and Equations Multiple Choice Questions 50 MCQs Nitrogen and Sulfur Multiple Choice Questions 89 MCQs Organic and Nitrogen Compounds Multiple Choice Questions 54 MCQs Periodicity Multiple Choice Questions 202 MCQs Polymerization Multiple Choice Questions 36 MCQs Rates of Reaction Multiple Choice Questions 39 MCQs Reaction Kinetics Multiple Choice Questions 52 MCQs Redox Reactions and Electrolysis Multiple Choice Questions 55 MCQs States of Matter Multiple Choice Questions 66 MCQs Transition Elements Multiple Choice Questions 29 MCQs The chapter Alcohols and Esters MCQs covers topics of introduction to alcohols and alcohols reactions The chapter Atomic Structure and Theory MCQs covers topics of atom facts elements and atoms number of nucleons protons electrons and neutrons The chapter Benzene Chemical Compound MCQs covers topics of benzene arenes reaction phenol properties and reactions of phenol The chapter Carbonyl Compounds MCQs covers topics of carbonyl compounds aldehydes and ketone testing nucleophilic addition with HCN

preparation of aldehydes and ketone reduction of aldehydes and ketone

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