



Time allowed: 2:40 Hours

Total Marks Sections B and C:

NOTE: Answer any eleven parts from Section 'B' and any two questions from Section 'C' on the separate provided answer book. Write your answers neatly and legibly. 09280208

SECTION - B (Marks 33)

(11 x 3 =

- Q. 2 Attempt any ELEVEN parts. All parts carry equal marks.
- (i) Briefly explain the following. Also give one example of each
 - (a) Ion
 - (b) Environmental Chemistry
 - (c) Atomic Mass
 - (ii) Compare compounds and mixtures. Also give examples.
 - (iii) There are three isotopes of hydrogen. Draw their diagrams.
 - (iv) Distinguish between 'shell' and 'subshell'.
 - (v) State periodic law. Also write the names of groups in s block of the periodic table.
 - (vi) Show the valence shell electronic configuration for the following:
 - a. Alkali metals ^{s¹}
 - b. Nobel gases ^{n², n²p⁶}
 - c. Halogens ^{n², n²p⁵}
 - (vii) Determine the location of C_{11}^{+} in periodic table. *G = P₂*
 - (viii) Briefly explain the following terms: (Any two)
 - a. Melting
 - b. Freezing
 - c. Boiling
 - (ix) Find the number of electrons in valence shell in the following:
 - a. Neon
 - b. Helium
 - c. Hydrogen
 - (x) Calculate the molarity of a solution that contains 40g urea (NH_2CONH_2) dissolved in 500cm³ of solution
14 + 2 + 2 + 12 + 16 + 14 + 16 + 16 = 82
 - (xi) The atomic number of an element is 20 and mass number is 40. How many Protons, Neutrons and Electrons does an atom of this element have?
 - (xii) Enlist the possible uses of an electrolytic cell.
 - (xiii) Find oxidation state of Nitrogen in any two of given compounds:
 - a. N_2O_5
 - b. HNO_3
 - c. NO_2
 - (xiv) Magnesium reacts with steam, but calcium reacts with cold water. Show this diversity in reactivity by chemical equations.
 - (xv) Why fluorine has the highest oxidizing power? Explain briefly.

SECTION - C (Marks 20)

(2 x 10 = 20)

- Note: Attempt any TWO questions. All questions carry equal marks.
- Q. 3 a. What is meant by evaporation? The water level in an aquarium decreases slowly, even though the tank doesn't leak. Interpret
- b. Compare the effects of Al_2O_3 and Fe_2O_3 formation on their parent metals. (*Al* and *Fe*)
- Q. 4 a. KCl and MgF_2 are two ionic compounds. Discuss the bond formation in each of these, Also draw electron dot and cross models.
- b. Briefly discuss the following types of solutions, also give one example of each:
- | | Solute | Solvent |
|-----|--------|---------|
| i. | Solids | Liquids |
| ii. | Solids | Solids |
- Q. 5 a. What are unsaturated, saturated and supersaturated solutions? Also give their comparison.
- b. Describe the formation of anions in Bromine and Phosphorous ions.

15A-I 2308 HA 09280208

SUPPLEMENTARY TABLE

Atomic No	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
Symbol	H	He	Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca
Mass No	1	4	7	9	11	12	14	16	19	20	23	24	27	28	31	32	35.5	40	39	40



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