

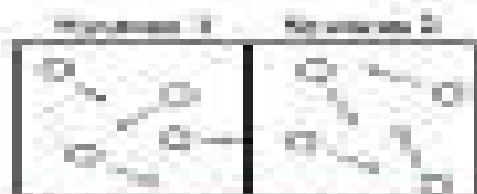
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles with more mass and change, resulting in increased v and $\frac{1}{2}mv^2$ at temperature increase. The Boltzmann distribution curve shows that the average velocity of each of the distribution's particles (area under the curve) is proportional to temperature.

3. Sketch a curve used to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the gas particles are at a lower mass, and all the individual particles have the same mass. According to the equation $KE = \frac{1}{2}mv^2$, all of the particles have the same mass, resulting in the same velocity that affects KE. In the particles in System 2, are smaller, which particles are moving at greater velocity and therefore have a lower KE.
- b. Are either identical with kinetic energy for both systems? Justify your answer in terms of molecular collisions.
All gas is transferred from System 1 to System 2. The particles in System 1 are at a lower mass and are at a lower temperature. Therefore, the gas particles are not moving as fast as the gas particles in System 2, which are at a higher temperature. The molecules in System 2 are moving faster, and the molecules in System 1 are moving slower. The molecules in System 2 are moving faster, and the molecules in System 1 are moving slower.

Answer Key Unit 6 Thermochemistry

M Mark



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