

SOLUBILITY CURVE WORKSHEET KEY

Use your solubility curve graphs provided to answer the following questions.

1. What are the customary units of solubility on solubility curves? Degress Celsius and grams of solute/100g of water
2. Define solubility. A measure of how much solute can dissolve in a given amount of solvent.
3. According to the graph, the solubility of any substance changes as temperature changes.
4. List the substances whose solubility decreases as temperature increases. NH_3 and $\text{Ce}_2(\text{SO}_4)_3$
5. Which substance is least affected by temperature changes? NaCl
6. How many grams of ammonium chloride (NH_4Cl) at 50°C ? 50g
7. NaCl and KClO_3 have the same solubility at approximately 78°C .
8. Which compound is least soluble in water at 10°C ? KClO_3
9. How many grams of KNO_3 can be dissolved at 50°C ? 80g
10. Are the following solutions unsaturated, saturated, or supersaturated?
 - a. 45g of NaNO_3 in 100 g of water at 30°C . saturated
 - b. 60g of KClO_3 in 100 g of water at 90°C . supersaturated
11. How many grams of sodium chloride, NaCl are required to saturate 100 grams of water at 100°C ? 40g
12. How many grams of NaNO_3 are required to saturate 100 grams of water at 75°C ? 140g
13. How many grams of KI will saturate water at 20°C ? 33g
14. At what temperature would 25g of potassium chlorate (KClO_3) dissolve? 60°C
15. At what temperature would 60g of NH_4Cl dissolve? 70°C
16. 89 g NaNO_3 is prepared at 30°C .
 - a) Will all of the salt dissolve? No
 - b) What mass of NaNO_3 will dissolve at this temperature? 95g
17. If 50 grams of NH_4Cl is dissolved at 50°C , how many additional grams NH_4Cl would be needed to make the solution saturated at 80°C ? 15g
18. At 50°C , how many grams of KNO_3 will dissolve? 80g
19. At 70°C , how many grams of cerium (III) sulfate ($\text{Ce}_2(\text{SO}_4)_3$) dissolve? 5g
20. Determine if each of the following is unsaturated, saturated, or supersaturated.
 - a. 55g of NH_3 at 20°C . supersaturated
 - b. 10g of $\text{Ce}_2(\text{SO}_4)_3$ at 10°C . unsaturated
 - c. 110g of KNO_3 at 60°C . supersaturated
 - d. 65g of NH_4Cl at 80°C . saturated
 - e. 12g of NH_3 at 90°C . supersaturated
 - f. 78g of NaNO_3 at 10°C . saturated
 - g. 145g of NaNO_3 at 80°C . saturated
 - h. 35g of NaCl at 100°C . unsaturated

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Quarterly Journal of the Chemical Society of London Chemical Society (Great Britain), 1958-04

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Atul,2010-09 *Chemical Engineering for Non-Chemical Engineers* Jack Hipple,2017-02-06 Outlines the concepts of chemical engineering so that non chemical engineers can interface with and understand basic chemical engineering concepts Overviews the difference between laboratory and industrial scale practice of chemistry consequences of mistakes and approaches needed to scale a lab reaction process to an operating scale Covers basics of chemical reaction engineering mass energy and fluid energy balances how economics are scaled and the nature of various types of flow sheets and how they are developed vs time of a project Details the basics of fluid flow and transport how fluid flow is characterized and explains the difference between positive displacement and centrifugal pumps along with their limitations and safety aspects of these differences Reviews the importance and approaches to controlling chemical processes and the safety aspects of controlling chemical processes Reviews the important chemical engineering design aspects of unit operations including distillation absorption and stripping adsorption evaporation and crystallization drying and solids handling polymer manufacture and the basics of tank and agitation system design **Educart CBSE Class 12 Chemistry One Shot Question Bank 2026**

(Includes PYQs for 2025-26) , Mass Transfer Operations Mr. Rohit Manglik,2024-07-24 Discusses theoretical and practical aspects of mass transfer including distillation absorption and drying for students of chemical engineering

Handbook of Industrial Crystallization Allan Myerson,2002-01-08 Crystallization is an important separation and purification process used in industries ranging from bulk commodity chemicals to specialty chemicals and pharmaceuticals In recent years a number of environmental applications have also come to rely on crystallization in waste treatment and recycling processes The authors provide an introduction to the field of newcomers and a reference to those involved in the various aspects of industrial crystallization It is a complete volume covering all aspects of industrial crystallization including material related to both fundamentals and applications This new edition presents detailed material on crystallization of biomolecules precipitation impurity crystal interactions solubility and design Provides an ideal introduction for industrial crystallization newcomers Serves as a worthwhile reference to anyone involved in the fieldCovers all aspects of industrial crystallization in a single complete volume

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