

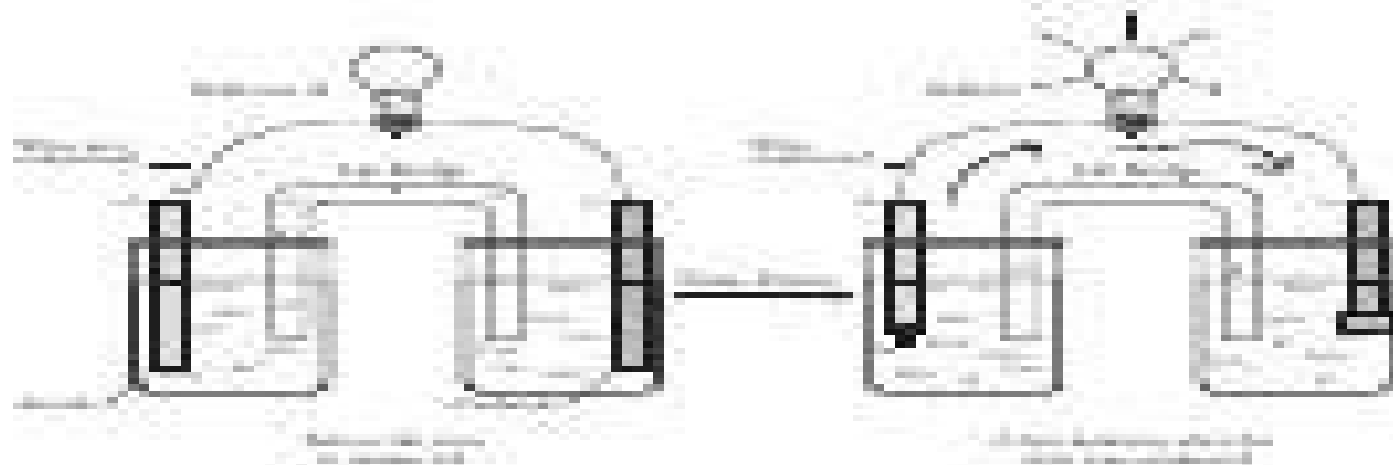
Discussion

(Work about 2-3 minutes; students will write)

Wrap-Up

Students will present their work and explain their answers. Students should be prepared to discuss why the two half-cells are not connected in the first half-cell. They should be able to explain why the two half-cells are connected in the second half-cell. They should be able to explain why the two half-cells are connected in the third half-cell. They should be able to explain why the two half-cells are connected in the fourth half-cell.

Student 1: Voltage Cell



1. Describe the structure of Student 1's voltaic cell. What are the two half-cells? What are the electrodes? What are the solutions? What is the salt bridge? What is the purpose of the salt bridge?

Students & teacher write down key points for discussion.

2. Describe the system in Student 2's voltaic cell. What are the two half-cells? What are the electrodes? What are the solutions? What is the salt bridge? What is the purpose of the salt bridge?

Students & teacher write down key points for discussion.

Answers Pogil How Does A Battery Work

Victoria G. Christensen



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The Book of Battery Manufacturing Razzab Ali Shaikh, 2014-10-06 What are the main parts of a battery All batteries contain one or more cells but people often use the terms battery and cell interchangeably A cell is just the working chemical unit inside a battery one battery can contain any number of cells A cell has three main parts a positive electrode terminal a negative electrode and a liquid or solid separating them called the electrolyte When a battery is connected to an electric circuit a chemical reaction takes place in the electrolyte causing ions in this case atoms with a positive electrical charge to flow through it one way with electrons particles with a negative charge flowing through the outer circuit in the other direction This movement of electric charge makes an electric current flow through the cell and through the circuit it is connected to That s the theory anyway Now let s look at it in practice How does a battery really work Where does the power in a battery actually come from Let s take a closer look Here s my battery hooked up to a flashlight bulb to make a simple circuit I ve unwrapped a paperclip to make a piece of connecting wire and I m holding that between the bottom of the battery and the side of the bulb If you look closely you can see the bulb is shining That s because electrons are marching through it Anode and cathode Now here s what s going on inside The battery s positive terminal shown just above my left thumb in the photo and colored red in the artwork below is connected to a positive electrode that s mostly hidden inside the battery We call this the cathode The outer case and the bottom of the battery make up the negative terminal or negative electrode which is also called the anode and colored green in the artwork The paperclip wire is represented in the art by the blue line Let s quickly clear up one point of confusion At school you may have learned that the cathode is the negative electrode and the anode the positive electrode However that

applies only to electrolysis passing electricity through a chemical to split it up Batteries are like electrolysis going backwards they split up chemicals to make electricity so the terms anode and cathode are switched around Okay To avoid confusion I suggest it's best not to use the terms anode and cathode at all It's better to say positive terminal and negative terminal and then it's always clear what you mean whether you're talking about batteries or electrolysis Chemical reactions Now back to our battery The positive and negative electrodes are separated by the chemical electrolyte It can be a liquid but in an ordinary battery it is more likely to be a dry powder When you connect the battery to a lamp and switch on chemical reactions start happening One of the reactions generates positive ions shown here as big yellow blobs and electrons smaller brown blobs at the negative electrode The positive ions flow through the electrolyte to the positive electrode from the green line to the red one Meanwhile the electrons smaller brown blobs flow around the outside circuit blue line to the positive electrode and make the lamp light up on the way The electrons and ions flow because of the chemical reactions happening inside the battery usually two or three of them going on simultaneously The exact reactions depend on the materials from which the electrodes and electrolyte are made and we won't go into them here If you want to know what they are Google the type of the battery you're interested in followed by the words anode cathode reactions For example Google zinc carbon battery anode cathode reactions Whatever chemical reactions take place the general principle of electrons going around the outer circuit and ions flowing in the opposite direction through the electrolyte happens in all batteries As the battery generates power the chemicals inside it are gradually converted into different chemicals

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