

6. Given the following half-reaction reduction of a galvanic cell:



- a. Identify the anode and cathode half-reactions:



Note: This problem is an iron electrode with iron ions only in solution.

- b. Write the overall balanced cell reaction:



6. Write the cell notation for a galvanic cell that uses the reactions below:



7. A galvanic cell is constructed based on the following reaction:



- a. Calculate the standard free energy change (ΔG°) for this reaction at 298 K.



$$= -2(96485 \text{ C/mol})E^{\circ}_{\text{cell}}$$

$$n = 2 \text{ mol e}^{-}$$



$$n = 2 \text{ mol e}^{-} \times (0.036 \text{ V}) \times (96485 \text{ C/mol}) = 7.4 \text{ kJ/mol}$$

$$n = 2(96485 \text{ C/mol})$$

$$n = 2(96485 \text{ C/mol})$$

- b. Is this reaction spontaneous?

Yes, this reaction is spontaneous at standard state conditions, as indicated by the negative sign of ΔG° .

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R Pring



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Textbook Problem Pack - Kimmel, Accounting 5e Kimmel, Paul D Kimmel, PhD, CPA, Jerry J. Weygandt, Donald E. Kieso, 2013-02-19 , *AECon 2020* Saefurrohman, Malim Muhammad, Heri Nurdiyanto , 2021-08-19 The 6th Asia Pacific Education and Science Conference AECON 2020 was conducted on 19-20 December 2020 at Universitas Muhammadiyah Purwokerto Purwokerto Indonesia The Theme of AECON 2020 is Empowering Human Development Through Science and Education The goals of AECON 2020 is to establish a paradigm that emphasizes on the development of integrated education and science through the integration of different life skills in order to improve the quality of human development in education and science around Asia Pacific nations particularly Indonesia *National Household Survey on Drug Abuse* , 2000

Survey Measurement of Drug Use , 1992 *Monthly Catalogue, United States Public Documents* , 1991 Pamphlet , 1952 **Catalog of the U.S. Military Academy** Military Academy, West Point, 1950 **Monthly Catalog of United States Government Publications** , 1995 **Specimen Questions from U.S. Civil Service Examinations** United States Civil Service Commission, 1952

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