energy level

CHAPTER 7 STUDY GUIDE

Ionic Compounds and Metals

Section 7.1 Ion Formation

chemical bond

In your textbook, read about chemical bonds and formation of ions.

electrons

Use each of the terms below just once to complete the passage.

becoming anions.

nucieus	ociet	pseudo-noble gas formations	valuence
The force	that holds two atoms together	is called a(n) (1)	
Such an attac	hment may form by the attract	ion of the positively charged	
(2)	of one atom for the negatively charged		
(3)	of another atom, or by the attraction of charged atoms, which are called		
	. The attractions may also involve		
(5)	electrons, whic	th are the electrons in the outermost	
(6)	. The (7)	are a family of elements	that have very
little tendenc	y to react. Most of these eleme	nts have a set of eight outermost electrons,	which is called a
stable (8)	. The relatively stable electron structures developed by loss of		
electrons in c	ertain elements of groups 3, 4,	13, and 14 are called (9)	
For each sta	tement below, write true or fa	alse.	
	10. A positively charged ion is called an anion.		
	 Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge. 		
	Elements tend to react so that they acquire the electron structure of a halogen.		
	 A sodium atom tends to lose one electron when it reacts. 		
	14. The electron structure of a zinc ion (Zn ²⁺) is an example of a pseudo-noble gas formation.		
	15. A Cl ⁻ ion is an examp	de of a cation.	
	16. The ending -ide is used to designate an anion.		
	17 Nonmostale form a sta-	ble outer electron configuration by losing a	destrone and

noble gases.

BOOK

Answer Study Guide Ionic Compounds And Metals

G Thomas

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AP Biology Study Guide Sundar Nathan, 2009-11 Sundar Nathan received a Bachelor's degree in Electrical Engineering from Anna University Chennai India and a Masters degree in Biomedical Engineering from the University of Texas at Austin Working for over a year with a team of talented Phds MPhils and MScs from all over the world Sundar compiled this comprehensive study guide to help students prepare diligently understand the concepts and Crush the AP Bio Test Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter "O" Level Study Guide - Chemistry Quite Easily Done,

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