



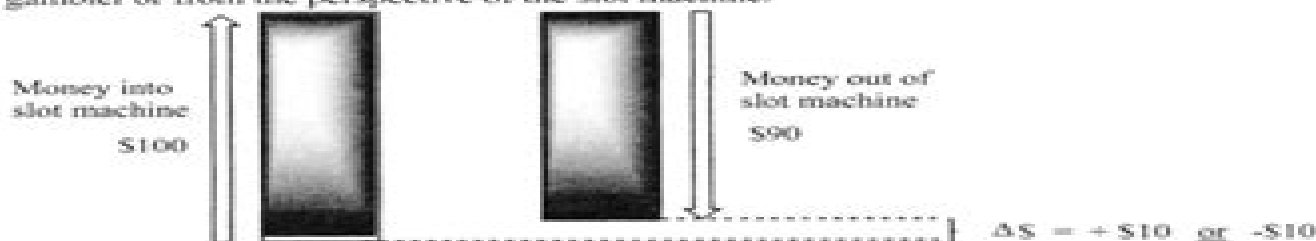
ChemActivity 6

Part A: Energy Diagrams

(What is the sign of $\Delta H_{\text{reaction}}$ for an exothermic reaction?)

Model 1A: The Slot Machine

Imagine a gambler sitting at a slot machine. Over the course of an hour the gambler puts \$100 into the machine and the machine pays out \$90. We know a net change (Δ) of ten dollars occurred, but what is the sign of this change in money (ΔS). To determine if $\Delta S = +10$ or -10 we need to decide if we calculate ΔS from the perspective of the gambler or from the perspective of the slot machine.



To translate this analogy into chemistry remember that:

- **Money = energy.**
- **The slot machine = the reaction.**
- **The gambler = the experimenter (you).**
- **All changes (Δ) will be calculated from the perspective of the reaction (a.k.a. the slot machine).**

Critical Thinking Questions

1. Using the convention listed above for calculating changes (Δ), what is the sign of ΔS in the example in Model 1A?
2. Based on this same convention, what sign must be attached to the \$100 that the gambler put into the slot machine in Model 1A? The sign of the \$90 that comes out is the opposite. **Add the appropriate sign to each value in Model 1A.**
3. On a different day the gambler puts in a total of \$20 and the slot machine pays out a total of \$100. What is the sign and value of ΔS for this day?



Answers Chemactivity 6

**Michael P. Garoutte, Ashley B.
Mahoney**



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