

Kinetics Worksheet 1

3. For the reaction: $A + 2B \rightarrow 3C + D$
the following data were obtained,

run	[A]	[B]	rate $M \cdot s^{-1}$
1	0.16	0.36	2.67
2	0.16	0.09	0.67
3	0.32	0.24	7.11
4	0.08	0.09	0.168
5	0.24	0.12	2.0
6	0.02	1.35	?

$$R = k[A]^y[B]^x$$

- (A) Write the rate law expression for this reaction.

(B) $\left(\frac{0.36}{0.09}\right)^x = \frac{2.67 \text{ m s}^{-1}}{0.67 \text{ m s}^{-1}}$ (A) $\left(\frac{0.16}{0.08}\right)^y = \frac{0.67 \text{ m s}^{-1}}{0.168 \text{ m s}^{-1}}$
 $(4)^x = 4$ $(2)^y = 4$
 $x \ln 4 = \ln 4$ $y \ln 2 = \ln 4$
 $x = 1$ $y = 2$

$$R = k[A]^2[B]$$

- (B) Calculate the value of k, the rate constant (remember units).

$$k = \frac{R}{[A]^2[B]} = \frac{2.67 \text{ m s}^{-1}}{(0.16)^2 [0.36]} \quad k = \text{m}^{-2} \text{s}^{-1}$$

$$k = 290 \text{ m}^{-2} \text{s}^{-1}$$

- (C) Calculate the rate of the reaction when $[A] = 0.02 \text{ M}$ and $[B] = 1.35 \text{ M}$.

$$R = 290 \text{ m}^{-2} \text{s}^{-1} [0.02]^2 [1.35] = 0.16 \frac{\text{m}}{\text{s}}$$

- (D) Assume that the reaction goes to completion. Under the conditions specified in the second run, what would the final [C] be?



$$\text{limit} \Rightarrow 0.09 B \times \frac{1A}{2B} = 0.045 A$$

$$0.045 A \times \frac{3C}{1A} = 0.135 \text{ m C}$$

$$.16 A \times \frac{3C}{1A} = .48 \text{ m}$$

$$.09 B \times \frac{3C}{2B} = .135 \text{ m}$$

or

Answers Kinetics Worksheet 1

Lilian W. Y. Cheung



Answers Kinetics Worksheet 1:

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Table of Contents Answers Kinetics Worksheet 1

1. Understanding the eBook Answers Kinetics Worksheet 1
 - The Rise of Digital Reading Answers Kinetics Worksheet 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Kinetics Worksheet 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Kinetics Worksheet 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Kinetics Worksheet 1
 - Personalized Recommendations

- Answers Kinetics Worksheet 1 User Reviews and Ratings
- Answers Kinetics Worksheet 1 and Bestseller Lists
- 5. Accessing Answers Kinetics Worksheet 1 Free and Paid eBooks
 - Answers Kinetics Worksheet 1 Public Domain eBooks
 - Answers Kinetics Worksheet 1 eBook Subscription Services
 - Answers Kinetics Worksheet 1 Budget-Friendly Options
- 6. Navigating Answers Kinetics Worksheet 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Kinetics Worksheet 1 Compatibility with Devices
 - Answers Kinetics Worksheet 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Kinetics Worksheet 1
 - Highlighting and Note-Taking Answers Kinetics Worksheet 1
 - Interactive Elements Answers Kinetics Worksheet 1
- 8. Staying Engaged with Answers Kinetics Worksheet 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Kinetics Worksheet 1
- 9. Balancing eBooks and Physical Books Answers Kinetics Worksheet 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Kinetics Worksheet 1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Kinetics Worksheet 1
 - Setting Reading Goals Answers Kinetics Worksheet 1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Kinetics Worksheet 1
 - Fact-Checking eBook Content of Answers Kinetics Worksheet 1

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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