

Name _____ answer key _____ Date _____

Stoichiometry: Mass – Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₃.

$$\frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3} \times \frac{1 \text{ mole FeCl}_3}{9 \text{ FeCl}_3} \times \frac{\text{mole NaCl}}{\text{mole FeCl}_3}$$

Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

$$\frac{9 \text{ Na}_2\text{SO}_4}{1 \text{ mole Na}_2\text{SO}_4} \times \frac{1 \text{ mole Na}_2\text{SO}_4}{9 \text{ Na}_2\text{SO}_4} \times \frac{\text{mole Fe}_2(\text{SO}_4)_3}{\text{mole Na}_2\text{SO}_4} \times \frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3}$$

Calculate the amount in grams of FeCl₃ produced with 95g of Fe (SO₄)₃

$$\frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3} \times \frac{1 \text{ mole Fe}_2(\text{SO}_4)_3}{9 \text{ Fe}_2(\text{SO}_4)_3} \times \frac{\text{mole FeCl}_3}{\text{mole Fe}(\text{SO}_4)_3} \times \frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3}$$

Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl

$$\frac{9 \text{ NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole NaCl}}{9 \text{ NaCl}} \times \frac{\text{mole Fe}(\text{SO}_4)_3}{\text{mole NaCl}}$$

Based on the following chemical equation, please answer the following questions!



$$\frac{9 \text{ Ca}(\text{HCO}_3)_2}{1 \text{ mole Ca}(\text{HCO}_3)_2} \times \frac{1 \text{ mole Ca}(\text{HCO}_3)_2}{9 \text{ Ca}(\text{HCO}_3)_2} \times \frac{\text{mole CaCO}_3}{\text{mole Ca}(\text{HCO}_3)_2}$$

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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. *Study Guide/Selected Solutions Manual* Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter, review of skills, self-tests, and the answers and detailed solutions to all odd-numbered end-of-chapter problems in the text book. *Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry* Susan McMurry, 1996 This internationally acclaimed detective series is just the thing for lovers of those Number One Ladies looking for a darker, more realistic view of Botswana. Sue Baker Publishing News. *Study Guide with Answers to Selected Problems* H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview, activities, and practice problems to reinforce skills, and a practice test. The answers section includes answers for all odd-numbered end-of-chapter exercises. **Study Guide and Full Solutions Manual** Susan McMurry, John McMurry, 2002-08 Contains a brief overview of every chapter, review of skills, self-tests, and the answers and detailed solutions to all end-of-chapter problems in the textbook. **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane, CH₄, from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston, Katherine J. Denniston, 2006 Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry Susan McMurry, 1992

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