## Advanced Placement Chemistry

#### 1995 Free Response Questions

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1) Solve the following problem.

 $H_2(g) + CO_2(g) \iff H_2O(g) + CO(g)$ 

When H<sub>2</sub>(g) is mixed with CO<sub>2</sub>(g) at 2,000 K, equilibrium is achieved according to the equation above. In one experiment, the following equilibrium concentrations were measured.

 $[H_2] = 0.20 \text{ mol/L}$   $[CO_2] = 0.30 \text{ mol/L}$  $[H_2O] = [CO] = 0.55 \text{ mol/L}$ 

(a) What is the mole fraction of CO(g) in the equilibrium mixture?

(b) Using the equilibrium concentrations given above, calculate the value of K<sub>e</sub>, the equilibrium constant for the reaction.

(c) Determine Kp, in terms of Kc for this system.

(d) When the system is cooled from 2,000 K to a lower temperature, 30.0 percent of the CO(g) is converted back to CO<sub>2</sub>(g). Calculate the value of K<sub>c</sub> at this lower temperature.

(e) In a different experiment, 0.50 mole of H<sub>2</sub>(g) is mixed with 0.50 mole of CO<sub>2</sub>(g) in a 3.0-liter reaction vessel at 2,000 K. Calculate the equilibrium concentration, in moles per liter, of CO(g) at this temperature.

Propane, C<sub>3</sub>H<sub>8</sub>, is a hydrocarbon that is commonly used as fuel for cooking.

(a) Write a balanced equation for the complete combustion of propane gas, which yields CO<sub>2</sub>(g) and H<sub>2</sub>O(I)

(b) Calculate the volume of air at 30°C and 1.00 atmosphere that is needed to burn completely 10.0 grams of propane. Assume that air is 21.0 percent O<sub>2</sub> by volume.

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**BM King** 

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