

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction:

This experiment involves the gravimetric analysis of a solution containing chloride ions. The procedure involves the precipitation of the chloride ions as silver chloride (AgCl) using a solution of silver nitrate (AgNO_3). The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to the toxic fumes of silver nitrate. The procedure is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until all the chloride ions have been precipitated as silver chloride. 3. Filtration: The precipitate is filtered from the solution using a filter paper. 4. Drying: The precipitate is dried in a desiccator to remove any water. 5. Weighing: The dried precipitate is weighed on a balance. 6. Calculation: The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

The reaction between silver nitrate and chloride ions is as follows: $\text{AgNO}_3 + \text{Cl}^- \rightarrow \text{AgCl} + \text{NO}_3^-$. The silver chloride precipitate is then filtered and dried. The mass of the dried precipitate is used to calculate the mass of chloride ions in the original solution. The calculation is as follows: $\text{Mass of Cl}^- = \frac{\text{Mass of AgCl}}{\text{Molar mass of AgCl}} \times \text{Molar mass of Cl}^-$.

Procedure:

The procedure for this experiment is as follows: 1. Preparation of the sample solution: A known mass of a sample containing chloride ions is dissolved in a known volume of water. 2. Addition of silver nitrate: A solution of silver nitrate is added to the sample solution until all the chloride ions have been precipitated as silver chloride. 3. Filtration: The precipitate is filtered from the solution using a filter paper. 4. Drying: The precipitate is dried in a desiccator to remove any water. 5. Weighing: The dried precipitate is weighed on a balance. 6. Calculation: The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

Data Analysis:

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride ion	0.350 g	0.350 g

The mass of chloride ion in the sample is calculated as follows: $\text{Mass of Cl}^- = \frac{\text{Mass of AgCl}}{\text{Molar mass of AgCl}} \times \text{Molar mass of Cl}^-$

For Sample 1: $\text{Mass of Cl}^- = \frac{0.850 \text{ g}}{143.32 \text{ g/mol}} \times 35.45 \text{ g/mol} = 0.350 \text{ g}$

For Sample 2: $\text{Mass of Cl}^- = \frac{0.850 \text{ g}}{143.32 \text{ g/mol}} \times 35.45 \text{ g/mol} = 0.350 \text{ g}$

Analysis Of Chloride Experiment 7 Lab Report

Peipei Pang



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