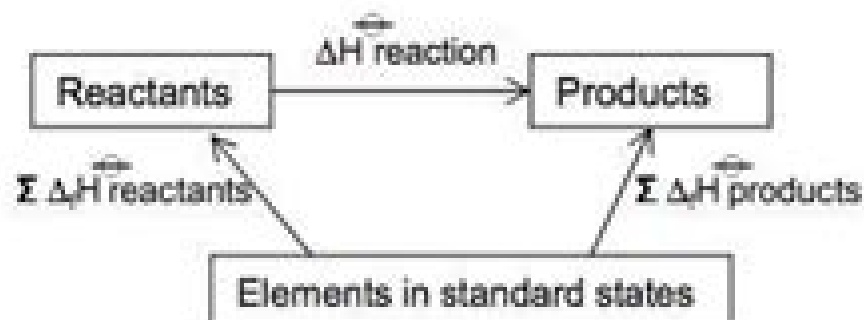


Using Hess's law to determine enthalpy changes from enthalpy changes of formation.

$$\Delta H_{\text{reaction}} = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$



Example 6. What is the enthalpy change for this reaction ?



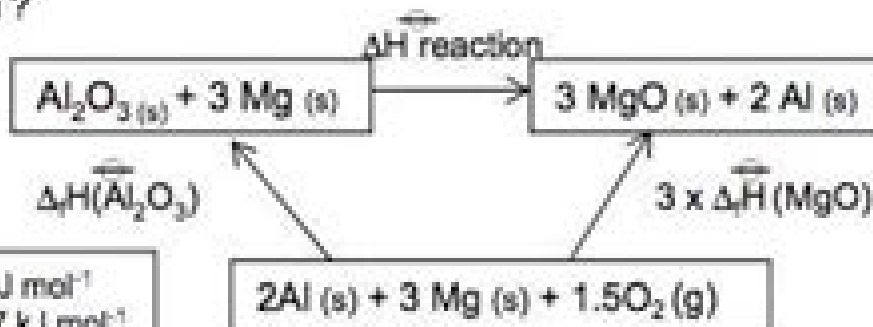
$$\Delta H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta H = 3 \times \Delta_f H (\text{MgO}) - \Delta_f H (\text{Al}_2\text{O}_3)$$

$$\begin{aligned} \Delta H &= (3 \times -601.7) - -1675.7 \\ &= -129.4 \text{ kJ mol}^{-1} \end{aligned}$$

Remember
elements have
 $\Delta_f H = 0$

$$\begin{aligned} \Delta_f H (\text{MgO}) &= -601.7 \text{ kJ mol}^{-1} \\ \Delta_f H (\text{Al}_2\text{O}_3) &= -1675.7 \text{ kJ mol}^{-1} \end{aligned}$$



Example 7. Using the following data to calculate the heat of combustion of propene

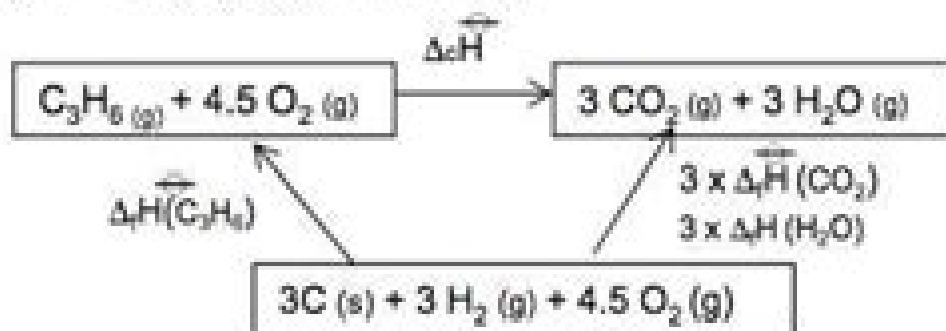
$$\Delta_f H, \text{C}_3\text{H}_6(\text{g}) = +20 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{CO}_2(\text{g}) = -394 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{H}_2\text{O}(\text{g}) = -242 \text{ kJ mol}^{-1}$$



$$\Delta_c H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta_c H = [3 \times \Delta_f H (\text{CO}_2) + 3 \times \Delta_f H (\text{H}_2\text{O})] - \Delta_f H (\text{C}_3\text{H}_6)$$

$$\begin{aligned} \Delta_c H &= [(3 \times -394) + (3 \times -242)] - 20 \\ &= -1928 \text{ kJ mol}^{-1} \end{aligned}$$



Answers For Chemfax Thermodynamics Enthalpy Lab

JS Bruner



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