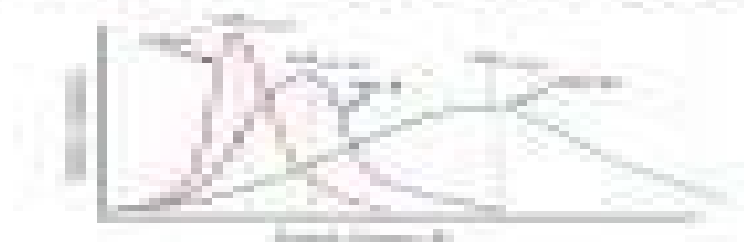


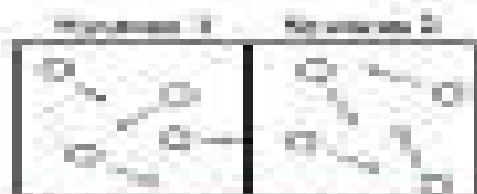
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving m/s is $\frac{1}{2}mv^2$, the $\frac{1}{2}$ is constant and m is constant, thus the change in kinetic energy can all be temperature increase. The Boltzmann distribution curve shows that the average velocity is all for the distribution's maximum value to change the average temperature increases.

3. Sketch a curve used to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more spread out, and all the individual particles have the same mass. According to the equation $\text{KE} = \frac{1}{2}mv^2$, all of the particles have the same mass, so the only way to get a higher KE is if the velocity is higher. In System 2, the particles are more spread out, and the average KE is lower, so the temperature is lower.
- b. Are there different total kinetic energy for System 1? Justify your answer in terms of molecular collisions.
No, the total kinetic energy is the same for both systems. In System 1, the particles are more spread out, and the average KE is lower, so the temperature is lower. In System 2, the particles are more spread out, and the average KE is higher, so the temperature is higher. The total kinetic energy is the same for both systems.

Answer Key Unit 6 Thermochemistry

Brendan G. Carr

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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