

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula of a compound without using percent forming (molar) masses of elements.
2. To calculate the number of moles of each of the reactants/products in the reaction.
3. To determine the percent reaction mass.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of silver nitrate: 5.00 g

Reaction of barium chloride: 5.00 g

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of silver nitrate solution: 5.00 g

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{moles of sodium chloride} = 25.0 \text{ mL sol. } \times \frac{1.0 \text{ g/mL}}{58.44 \text{ g/mol}} = 0.428 \text{ mol NaCl}$$

$$\text{moles of calcium chloride} = 25.0 \text{ mL sol. } \times \frac{1.0 \text{ g/mL}}{110.98 \text{ g/mol}} = 0.225 \text{ mol CaCl}_2$$

$$\text{moles of silver nitrate produced} = 5.00 \text{ g} \times \frac{1 \text{ mol}}{169.87 \text{ g/mol}} = 0.0295 \text{ mol AgNO}_3$$

$$\text{moles of product that would be produced} = 5.00 \text{ g CaCl}_2 \times \frac{1 \text{ mol}}{110.98 \text{ g/mol}} = 0.0451 \text{ mol CaCl}_2$$

2. Chemical equation: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Reaction:

$$\text{NaCl} + \text{CaCl}_2 \rightarrow \text{CaCl}_2 + 2\text{NaCl} \quad \text{This reaction is not a reaction.}$$

$$\text{NaCl} + \text{CaCl}_2 \rightarrow \text{CaCl}_2 + 2\text{NaCl}$$

$$\text{moles of CaCl}_2 \text{ formed in the reaction} = 0.0451 \text{ mol CaCl}_2 \times \frac{1 \text{ mol CaCl}_2}{110.98 \text{ g/mol}} = 0.00407 \text{ mol CaCl}_2$$

Answers For Lab 6b Heath Chemistry

Ulysses Orange Cox



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