

Chemistry

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(Chapter 1)(Solutions) (Intext Questions)

(Class 12)

Question 1.1:

Calculate the mass percentage of benzene (C_6H_6) and carbon tetrachloride (CCl_4) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride.

Answer 1.1:

$$\text{Mass percentage of } C_6H_6 = \frac{\text{Mass of } C_6H_6}{\text{Total mass of the solutions}} \times 100 \%$$

$$= \frac{\text{Mass of } C_6H_6}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{22}{22 + 122} \times 100 \% = 15.28 \%$$

$$\text{Mass percentage of } CCl_4 = \frac{\text{Mass of } CCl_4}{\text{Total mass of the solutions}} \times 100 \%$$

$$= \frac{\text{Mass of } CCl_4}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{122}{22 + 122} \times 100 \% = 84.72 \%$$

$$\text{Alternatively, Mass percentage of } CCl_4 = (100 - 15.28) \% = 84.72 \%$$

Question 1.2:

Calculate the mole fraction of benzene in solution containing 30% by mass in carbon tetrachloride.

Answer 1.2:

Let the total mass of the solution be 100 g and the mass of benzene be 30 g.

$$\therefore \text{Mass of carbon tetrachloride} = (100 - 30) \text{ g} = 70 \text{ g}$$

$$\text{Molar mass of benzene } (C_6H_6) = (6 \times 12 + 6 \times 1) \text{ g mol}^{-1} = 78 \text{ g mol}^{-1}$$

$$\therefore \text{Number of moles of } C_6H_6 = \frac{30}{78} \text{ mol} = 0.3846 \text{ mol}$$

$$\text{Molar mass of carbon tetrachloride } (CCl_4) = 1 \times 12 + 4 \times 35.5 \\ = 154 \text{ g mol}^{-1}$$

$$\therefore \text{Number of moles of } CCl_4 = \frac{70}{154} \text{ mol} = 0.4545 \text{ mol}$$

Thus, the mole fraction of C_6H_6 is given as:

$$= \frac{\text{Number of moles of } C_6H_6}{\text{Number of moles of } C_6H_6 + \text{Number of moles of } CCl_4}$$

$$= \frac{0.3846}{0.3846 + 0.4545}$$

$$= 0.458$$

Question 1.3:

Calculate the molarity of each of the following solutions:

(a) 30 g of $Co(NO_3)_2 \cdot 6H_2O$ in 4.3 L of solution

(b) 30 mL of 0.5 M H_2SO_4 diluted to 500 mL

Answer 1.3:

$$\text{Molarity is given by: } \text{Molarity} = \frac{\text{Moles of solute}}{\text{Volume of solution in litre}}$$

$$\text{(a) Molar mass of } Co(NO_3)_2 \cdot 6H_2O = 59 + 2(14 + 3 \times 16) + 6 \times 18 = 291 \text{ g mol}^{-1}$$

$$\therefore \text{Moles of } Co(NO_3)_2 \cdot 6H_2O = \frac{30}{291} \text{ mol} = 0.103 \text{ mol}$$

$$\text{Therefore, molarity} = \frac{0.103 \text{ mol}}{4.3 \text{ L}} = 0.023 \text{ M}$$

$$\text{(b) Number of moles present in 1000 mL of 0.5 M } H_2SO_4 = 0.5 \text{ mol}$$

$$\therefore \text{Number of moles present in 30 mL of 0.5 M } H_2SO_4 = \frac{0.5 \times 30}{1000} \text{ mol} = 0.015 \text{ mol}$$

$$\text{Therefore, molarity} = \frac{0.015 \text{ mol}}{0.5 \text{ L}} = 0.03 \text{ M}$$

8 Chemistry Sol Answers

**Richard Post, Chad Snyder, Clifford C.
Houk**



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Solutions Manual to Accompany Elements of Physical Chemistry David Smith, 2013-05-30 The Solutions Manual to accompany Elements of Physical Chemistry 6th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions

Solutions Manual to Accompany Physical Chemistry for the Life Sciences C. A. Trapp, M. P. Cady, 2011 This solutions manual contains fully worked solutions to all end of chapter discussion questions and exercises featured in Physical Chemistry for the Life Sciences

Longman Science Chemistry 9 Kohli Nitin, 2008-09 *Chapter-wise NCERT + Exemplar + Practice Questions with Solutions for CBSE Chemistry Class 11 - 2nd Edition* Disha Experts, 2017-08-29 The book Chapter wise NCERT Exemplar Practice Questions with Solutions for CBSE Class 11 Chemistry has been divided into 3 parts Part A provides detailed solutions Question by Question of all the questions exercises provided in the NCERT Textbook Part B provides solutions to the questions in the NCERT Exemplar book Part C provides selected Practice Questions useful for the Class 11 examination along with detailed solutions The solutions have been designed in such a manner Step by Step that it would bring 100% Concept Clarity for the student

Excel With Objective Chemistry For Iit-Screening Prof. S. K. Khanna, Dr. N. K. Sharma, Dr. B. Kapila, 2006

The Chemical News and Journal of Physical Science, 1893

A Reader's Guide to Contemporary Literature William Swan Sonnenschein, 1895

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