

MIXED GAS LAWS WORKSHEET - SOLUTIONS

1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

$$n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ Latm/mol.K})(292 \text{ K})} = 11 \text{ moles of gas}$$

2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of $25^\circ C$, what will the pressure of the resulting mixture of gases be? $25^\circ C = 298 \text{ K}$

$$O_2: P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$$

$$N_2: P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$$

$$P_{\text{tot}} = P_{O_2} + P_{N_2} = 4.1 \text{ atm} + 2.4 \text{ atm} = 6.5 \text{ atm}$$

Alternately you can substitute total moles of gas (8.0) as the total pressure depends on the total moles of gas – what type of gas is irrelevant

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is $20.0^\circ C$. By noon the temperature has risen to $45.0^\circ C$. What is the new volume of the balloon?

$$T_1 = 20.0^\circ C = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^\circ C = 318 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

$$n = \frac{PV}{RT} = \frac{(190 \text{ atm})(35 \text{ L})}{(0.0821 \text{ Latm/mol.K})(315 \text{ K})} = 260 \text{ moles of gas}$$

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in $^\circ C$ of the balloon?

$$T = \frac{PV}{nR} = \frac{(1 \text{ atm})(85 \text{ L})}{(3.5 \text{ mol})(0.0821 \text{ Latm/mol.K})} = 296 \text{ K} = 23^\circ C$$

6) $CaCO_3$ decomposes at $1200^\circ C$ to form CO_2 gas and CaO . If 25 L of CO_2 are collected at $1200^\circ C$, what will the volume of this gas be after it cools to $25^\circ C$?

$$T_1 = 1200^\circ C = 1473 \text{ K}, V_1 = 25 \text{ L}, T_2 = 25^\circ C = 298 \text{ K}, V_2 = ?$$

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at $20.0^\circ C$ is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is $-20.0^\circ C$?

$$P_1 = 1.00 \text{ atm}, V_1 = 4.50 \text{ L}, T_1 = 20.0^\circ C = 293 \text{ K}, P_2 = 0.600 \text{ atm}, V_2 = ?, T_2 = -20.0^\circ C = 253 \text{ K}$$

$$V_2 = \frac{P_1 V_1 T_2}{T_1 P_2} = \frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L}$$

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